

Chapter 17: Chemical Equilibrium

Worksheet: K<sub>c</sub> and K<sub>p</sub>

1. For the reaction  $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$

- 1. The equilibrium constant  $K_c$  for the reaction  $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$  is  $1.6 \times 10^3$  at  $1000^\circ\text{C}$ . Calculate  $K_c$  for the reaction  $\text{SO}_2(g) + \frac{1}{2}\text{O}_2(g) \rightleftharpoons \text{SO}_3(g)$  at  $1000^\circ\text{C}$ .
- 2. For the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$ ,  $K_c = 0.50$  at  $400^\circ\text{C}$ . Calculate  $K_c$  for the reaction  $2\text{NH}_3(g) \rightleftharpoons \text{N}_2(g) + 3\text{H}_2(g)$  at  $400^\circ\text{C}$ .
- 3. For the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$ ,  $K_c = 0.50$  at  $400^\circ\text{C}$ . Calculate  $K_c$  for the reaction  $\text{NH}_3(g) \rightleftharpoons \frac{1}{2}\text{N}_2(g) + \frac{3}{2}\text{H}_2(g)$  at  $400^\circ\text{C}$ .
- 4. For the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$ ,  $K_c = 0.50$  at  $400^\circ\text{C}$ . Calculate  $K_c$  for the reaction  $2\text{NH}_3(g) \rightleftharpoons \text{N}_2(g) + 3\text{H}_2(g)$  at  $400^\circ\text{C}$ .
- 5. For the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$ ,  $K_c = 0.50$  at  $400^\circ\text{C}$ . Calculate  $K_c$  for the reaction  $\text{NH}_3(g) \rightleftharpoons \frac{1}{2}\text{N}_2(g) + \frac{3}{2}\text{H}_2(g)$  at  $400^\circ\text{C}$ .
- 6. For the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$ ,  $K_c = 0.50$  at  $400^\circ\text{C}$ . Calculate  $K_c$  for the reaction  $2\text{NH}_3(g) \rightleftharpoons \text{N}_2(g) + 3\text{H}_2(g)$  at  $400^\circ\text{C}$ .
- 7. For the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$ ,  $K_c = 0.50$  at  $400^\circ\text{C}$ . Calculate  $K_c$  for the reaction  $\text{NH}_3(g) \rightleftharpoons \frac{1}{2}\text{N}_2(g) + \frac{3}{2}\text{H}_2(g)$  at  $400^\circ\text{C}$ .
- 8. For the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$ ,  $K_c = 0.50$  at  $400^\circ\text{C}$ . Calculate  $K_c$  for the reaction  $2\text{NH}_3(g) \rightleftharpoons \text{N}_2(g) + 3\text{H}_2(g)$  at  $400^\circ\text{C}$ .
- 9. For the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$ ,  $K_c = 0.50$  at  $400^\circ\text{C}$ . Calculate  $K_c$  for the reaction  $\text{NH}_3(g) \rightleftharpoons \frac{1}{2}\text{N}_2(g) + \frac{3}{2}\text{H}_2(g)$  at  $400^\circ\text{C}$ .
- 10. For the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$ ,  $K_c = 0.50$  at  $400^\circ\text{C}$ . Calculate  $K_c$  for the reaction  $2\text{NH}_3(g) \rightleftharpoons \text{N}_2(g) + 3\text{H}_2(g)$  at  $400^\circ\text{C}$ .

Answers

- 1.  $1.6 \times 10^3$
- 2.  $0.50$
- 3.  $0.50$
- 4.  $0.50$
- 5.  $0.50$
- 6.  $0.50$
- 7.  $0.50$
- 8.  $0.50$
- 9.  $0.50$
- 10.  $0.50$

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**J Rink**



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