

## AP Chemistry Chapter 9 Answers – Zumdahl

### 9.15

$\text{H}_2\text{O}$  has  $2(1) + 6 = 8$  valance electrons

$\text{H}_2\text{O}$  has a tetrahedral arrangement of the electron pairs about the O atom that requires  $sp^3$  hybridization. Two of the four  $sp^3$  hybrid orbitals are used to form bonds to the two hydrogen atoms and the other two  $sp^3$  hybrid orbitals hold the two lone pairs of oxygen. The two O-H bonds are formed from overlap of the  $sp^3$  hybrid orbitals on oxygen with the 1s atomic orbitals on the hydrogen atoms.

### 9.17

$\text{H}_2\text{CO}$  has  $2(1) + 4 + 6 = 12$  valence electrons

The central carbon atom has a trigonal planar arrangement of the electron pairs which requires  $sp^2$  hybridization. The two C-H sigma bonds are formed from overlap of the  $sp^2$  hybrid orbitals on carbon with the hydrogen 1s atomic orbitals. The double bond between carbon and oxygen consists of one  $\sigma$  and one  $\pi$  bond. The oxygen atom, like the carbon atom, also has a trigonal planar arrangement of the electrons which requires  $sp^2$  hybridization. The  $\sigma$  bond in the double bond is formed from overlap of a carbon  $sp^2$  hybrid orbital with an oxygen  $sp^2$  hybrid orbital. The  $\pi$  bond in the double bond is formed from overlap of the unhybridized p atomic orbitals. Carbon and oxygen each have one unhybridized p atomic orbital which are parallel to each other. When two parallel p atomic orbitals overlap, a  $\pi$  bond results.

### 9.19

Ethane,  $\text{C}_2\text{H}_6$ , has  $2(4) + 6(1) = 14$  valence electrons.

The carbon atoms are  $sp^3$  hybridized. The six C-H sigma bonds are formed from overlap of the  $sp^3$  hybrid orbitals on C with the 1s atomic orbitals from the hydrogen atoms. The carbon-carbon sigma bond is formed from overlap of an  $sp^3$  hybrid orbital on each C atom.

Ethanol,  $\text{C}_2\text{H}_5\text{O}$  has  $2(4) + 6(1) + 6 = 20 e^-$

The two C atoms and the O atom are  $sp^3$  hybridized. All bonds are formed from overlap with these  $sp^3$  hybrid orbitals. The C-H and O-H sigma bonds are formed from overlap of  $sp^3$  hybrid orbitals with hydrogen 1s atomic orbitals. The C-C and C-O sigma bonds are formed from overlap of the  $sp^3$  hybrid orbitals on each atom.

### 9.21

See Exercises 8.67 and 8.73 for the Lewis structures. To predict the hybridization, first determine the arrangement of electron pairs about each central atom using the VSEPR model; then utilize the information in Figure 9.24 of the text to deduce the hybridization required for that arrangement of electron pairs.

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**Thomas Griffiths**



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