

Table 8.3 Partial Orbital Diagrams and Electron Configurations* for the Elements in Period 3

Atomic Number/ Element	Partial Orbital Diagram (3s and 3p Sublevels Only)		Full Electron Configuration	Condensed Electron Configuration
11/Na	3s ↑	3p □ □ □	$[1s^2 2s^2 2p^6] 3s^1$	[Ne] $3s^1$
12/Mg	↑↓	□ □ □	$[1s^2 2s^2 2p^6] 3s^2$	[Ne] $3s^2$
13/Al	↑↓	↑ □ □	$[1s^2 2s^2 2p^6] 3s^2 3p^1$	[Ne] $3s^2 3p^1$
14/Si	↑↓	↑ ↑ □	$[1s^2 2s^2 2p^6] 3s^2 3p^2$	[Ne] $3s^2 3p^2$
15/P	↑↓	↑ ↑ ↑	$[1s^2 2s^2 2p^6] 3s^2 3p^3$	[Ne] $3s^2 3p^3$
16/S	↑↓	↑↓ ↑ ↑	$[1s^2 2s^2 2p^6] 3s^2 3p^4$	[Ne] $3s^2 3p^4$
17/Cl	↑↓	↑↓ ↑↓ ↑	$[1s^2 2s^2 2p^6] 3s^2 3p^5$	[Ne] $3s^2 3p^5$
18/Ar	↑↓	↑↓ ↑↓ ↑↓	$[1s^2 2s^2 2p^6] 3s^2 3p^6$	[Ne] $3s^2 3p^6$

Hund's rule

Chapter 8 Electron Configuration And General Chemistry

**Darrell D. Ebbing, Larry K.
Krannich, Joan I. Senyk**



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