

Chemical Equations and Reactions

QUIZ
CH #8
SECTION #1

Section Quiz: Describing Chemical Reactions

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- B 1. Which of the following is true regarding the law of conservation of mass?
- It dictates that the number of molecules on each side of a chemical equation must be the same.
 - It dictates that the number of atoms of each element must be the same on both sides of a chemical equation.
 - It states that the mass of the reactants must remain constant in order for a chemical reaction to proceed.
 - It does not apply to chemical reactions.
- D 2. Which of the following may indicate that a chemical reaction has occurred?
- release of energy as light
 - a color change
 - gas bubble formation
 - All of the above
- D 3. Which of the following is a product in the reaction described by the word equation below?
- $$\text{iron} + \text{copper(II) sulfate} \rightarrow \text{iron(II) sulfate} + \text{copper}$$
- iron(II) sulfate
 - copper
 - iron
 - Both (a) and (b)
- C 4. What does CH_4 represent in the balanced chemical equation shown below?
- $$\text{C} + 2\text{H}_2 \rightarrow \text{CH}_4$$
- one atom of CH_4
 - five atoms of CH_4
 - one molecule of CH_4
 - 1 g of CH_4
- C 5. Which of the following is not true regarding a properly written chemical equation?
- It must be balanced.
 - It must contain proper chemical formulas.
 - It must show the formation of a precipitate or a gas.
 - It should represent known data.

Chemical Reactions Chapter Test Answer Key

Joacim Rocklöv



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Laboratory Chemistry Richard Bishop Moore, 1904 An Intermediate Textbook of Physiological Chemistry with Experiments Chauncey John Vallette Pettibone, 1917 Covers the general field of physiological chemistry in such a way as to give students a familiarity with compounds important from a biochemical viewpoint and to acquaint them with the fundamental processes which go on in the animal body Preface **Laboratory Directions for Elementary Chemistry** Helen Isham Mattill, 1914 Physical Chemistry and Biophysics for Students of Biology and Medicine Matthew Steel, 1928
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