

Chemistry - Chapter 7 Study Guide

Section 7.1:

- *Valence electrons* – electrons in the outer energy level
 - responsible for bonding with other atoms
- How do you know how many valence electrons an element has?
 - Look at the group number
 - Group 1A = 1 valence electron
 - Group 7A = 7 valence electrons
 - Etc...
- *Electron Dot Diagrams* – a way of showing and keeping track of valence electrons
 - Written using the following steps:
 1. Write the element symbol – representing the nucleus and inner electrons
 2. Put one dot for each valence electrons,
 3. Don't forget to add them using Hund's rule – they don't pair up unless they have to.
- *The Octet Rule* – atoms tend to want to achieve a noble gas configuration because it is the most stable. Octet = eight electrons
- *Cations* – a positive ion (usually metals)
 - Lose electrons to attain a noble gas configuration
 - Written like Na^{1+}
 - No dots are shown (on an electron dot diagram) because they lose the valence electrons they had
- *Anions* – a negative ion (usually non-metals)
 - Gain electrons to attain a noble gas configuration
 - Written like P^{3-}
 - Dots are shown because valence electrons are gained
- 8 valence electrons = stable atom (noble gases.)

Section 7.2:

- Anions (-) and cations (+) are held together by opposite charges; a bond is formed through the transfer of electrons to achieve a noble gas configuration (octet.)
- *Ionic Compounds:*
 - Also called salts
 - Made from a cation and an anion
 - Which means they're made of a metal and a non-metal
- ***if you don't understand how to do ionic bonding, check the book or ask somebody to help... it's hard to explain***
- The result of ionic bonding is an ionic compound, shown in a *formula unit*
 - Written like Ca_3P_2

Chemistry Chapter 7 Study Guide Answers

Jin-Ying Zhang



Chemistry Chapter 7 Study Guide Answers:

Embracing the Song of Term: An Emotional Symphony within **Chemistry Chapter 7 Study Guide Answers**

In a world consumed by displays and the ceaseless chatter of quick connection, the melodic splendor and emotional symphony created by the prepared word often disappear into the back ground, eclipsed by the relentless sound and distractions that permeate our lives. But, situated within the pages of **Chemistry Chapter 7 Study Guide Answers** a marvelous fictional value filled with organic thoughts, lies an immersive symphony waiting to be embraced. Crafted by an elegant musician of language, that captivating masterpiece conducts viewers on a mental journey, well unraveling the hidden tunes and profound impact resonating within each cautiously crafted phrase. Within the depths of the moving examination, we can discover the book is main harmonies, analyze its enthralling writing design, and submit ourselves to the profound resonance that echoes in the depths of readers souls.

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