

Chemistry

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(Chapter 1)(Solutions) (Intext Questions)
(Class 12)

Question 1.11:

Calculate the mass of ascorbic acid (Vitamin C, $C_6H_8O_6$) to be dissolved in 75 g of acetic acid to lower its melting point by 1.5°C . $K_f = 3.9 \text{ K kg mol}^{-1}$.

Answer 1.11:

Mass of acetic acid, $w_1 = 75 \text{ g}$

Molar mass of ascorbic acid ($C_6H_8O_6$), $M_2 = 6 \times 12 + 8 \times 1 + 6 \times 16 = 176 \text{ g mol}^{-1}$

Lowering of melting point, $\Delta T_f = 1.5 \text{ K}$

We know that:

$$\Delta T_f = \frac{K_f \times 1000 \times w_2}{M_2 \times w_1}$$

$$\Rightarrow w_2 = \frac{\Delta T_f \times M_2 \times w_1}{K_f \times 1000}$$

$$= \frac{0.5 \times 176 \times 75}{3.9 \times 1000}$$

$$= 5.08 \text{ g (approx)}$$

Hence, 5.08 g of ascorbic acid is needed to be dissolved.

Question 1.12:

Calculate the osmotic pressure in Pascals exerted by a solution prepared by dissolving 1.0 g of polymer of molar mass 185,000 in 450 mL of water at 37°C .

Answer 1.12:

It is given that:

Volume of water, $V = 450 \text{ mL} = 0.45 \text{ L}$

Temperature, $T = (37 + 273) \text{ K} = 310 \text{ K}$

Number of moles of the polymer, $n = \frac{1}{185000} \text{ mol}$

We know that: $\pi = \frac{n}{V} RT$

Osmotic pressure,

$$= \frac{1}{185000} \text{ mol} \times \frac{1}{0.45 \text{ L}} \times 8.314 \times 10^3 \text{ Pa L K}^{-1} \times 310 \text{ K}$$

$$= 30.98 \text{ Pa}$$

$$= 31 \text{ Pa (approximately)}$$

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Chemistry Chapter 12 Solutions Answers

AW Rasmussen



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Table of Contents Chemistry Chapter 12 Solutions Answers

1. Understanding the eBook Chemistry Chapter 12 Solutions Answers
 - The Rise of Digital Reading Chemistry Chapter 12 Solutions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 12 Solutions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 12 Solutions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 12 Solutions Answers
 - Personalized Recommendations
 - Chemistry Chapter 12 Solutions Answers User Reviews and Ratings
 - Chemistry Chapter 12 Solutions Answers and Bestseller Lists

5. Accessing Chemistry Chapter 12 Solutions Answers Free and Paid eBooks
 - Chemistry Chapter 12 Solutions Answers Public Domain eBooks
 - Chemistry Chapter 12 Solutions Answers eBook Subscription Services
 - Chemistry Chapter 12 Solutions Answers Budget-Friendly Options
6. Navigating Chemistry Chapter 12 Solutions Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 12 Solutions Answers Compatibility with Devices
 - Chemistry Chapter 12 Solutions Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 12 Solutions Answers
 - Highlighting and Note-Taking Chemistry Chapter 12 Solutions Answers
 - Interactive Elements Chemistry Chapter 12 Solutions Answers
8. Staying Engaged with Chemistry Chapter 12 Solutions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 12 Solutions Answers
9. Balancing eBooks and Physical Books Chemistry Chapter 12 Solutions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 12 Solutions Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 12 Solutions Answers
 - Setting Reading Goals Chemistry Chapter 12 Solutions Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 12 Solutions Answers
 - Fact-Checking eBook Content of Chemistry Chapter 12 Solutions Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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