

## Chapter Test II

### A. Matching

Match each item in Column I with the correct description in Column II. Write the letter of the correct description.

#### Column I

- \_\_\_\_\_ 1. An unbalanced chemical equation for a reaction that obeys the law of conservation of matter.
- \_\_\_\_\_ 2. The number of protons in the nucleus of an atom of an element.
- \_\_\_\_\_ 3. The number of electrons in a neutral atom.
- \_\_\_\_\_ 4. The ratio of the actual weight of an element to the standard weight of that element.
- \_\_\_\_\_ 5. The number of protons and neutrons in the nucleus of an atom.
- \_\_\_\_\_ 6. An element whose atoms have a mass number of 12.

#### Column II

- a. atomic weight
- b. limiting reagent
- c. molecular weight
- d. molar mass
- e. atomic number
- f. atomic weight

### B. Multiple Choice

Choose the best answer for each question.

- \_\_\_\_\_ 1. How many molecules of  $\text{CO}_2$  are produced when 100 g of  $\text{C}_2\text{H}_6$  reacts with 100 g of  $\text{O}_2$ ?
 

<ol style="list-style-type: none"> <li>a. 1.00 mole <math>\text{CO}_2</math></li> <li>b. 1.00 mole <math>\text{CO}_2</math></li> <li>c. 1.00 mole <math>\text{CO}_2</math></li> <li>d. 1.00 mole <math>\text{CO}_2</math></li> </ol>	<ol style="list-style-type: none"> <li>e. 1.00 mole <math>\text{CO}_2</math></li> <li>f. 1.00 mole <math>\text{CO}_2</math></li> <li>g. 1.00 mole <math>\text{CO}_2</math></li> <li>h. 1.00 mole <math>\text{CO}_2</math></li> </ol>
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- \_\_\_\_\_ 2. Which of the following is not a chemical reaction?
  - a.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
  - b.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$
  - c.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{O}_2$
  - d.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{O}_2$
- \_\_\_\_\_ 3. How many grams of  $\text{O}_2$  are produced when 100 g of  $\text{KClO}_3$  is heated to produce  $\text{KCl}$  and  $\text{O}_2$ ?
 

<ol style="list-style-type: none"> <li>a. 100 g</li> <li>b. 100 g</li> </ol>	<ol style="list-style-type: none"> <li>c. 100 g</li> <li>d. 100 g</li> </ol>
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# Chemistry Chapter 12 Stoichiometry Test

**Mary Millhollon, Richard H. Langley**



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tables and a glossary of chemistry terms

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