

Grade 12 Chemistry Exam Review

Below is a list of major topics covered on the exam for this course. If a topic is not listed, it will NOT be on the exam.

Your studying should begin with your class notes, assigned problems, textbook problems and extra review questions suggested below. This list is a guide and some specific questions on the exam may not appear here.

Good luck ☺

Unit #1 – Energy Changes Text Reference: Chapter 5

- Energy changes in reactions (system, surroundings, endothermic vs. exothermic)
- Specific Heat Capacity
- Calorimetry, $q=mc\Delta T$
- Molar enthalpies
- Representing Enthalpy Changes (Thermochemical Equations, Potential Energy Diagrams, etc.)
- Hess' Law
- Standard Enthalpies of Formation

Sample Exam Questions

- Given the following information:

$$2\text{CH}_3\text{OH}(\text{l}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 4\text{H}_2\text{O}(\text{l}) \quad \Delta H^\circ = -727 \text{ kJ}\cdot\text{mol}^{-1}$$

$$\Delta H_f^\circ \text{CO}_2(\text{g}) = -394 \text{ kJ}\cdot\text{mol}^{-1}, \quad \Delta H_f^\circ \text{H}_2\text{O}(\text{l}) = -286 \text{ kJ}\cdot\text{mol}^{-1}$$

What is the standard heat formation, ΔH_f° (kJ·mol⁻¹) of CH₃OH(l)?

A. -602 B. -1329 C. 1329 D. 602
- When 789 g of liquid ammonia is cooled from 82.7 °C to 25.0 °C it loses 214 kJ of heat. What is the specific heat capacity of liquid ammonia?

A. 0.213 B. 4.19 C. 0.0047 D. 2.06 E. 4.70
- Identify the correct expression that can be used to determine the change in thermal energy for the following reaction:

$$\text{C}_2\text{H}_2(\text{g}) + \frac{13}{2}\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 5\text{H}_2\text{O}(\text{l}) + \text{heat?}$$

A. $\Delta H^\circ = [(4 \text{ mol})(\Delta H_f^\circ \text{CO}_2(\text{g})) + (5 \text{ mol})(\Delta H_f^\circ \text{H}_2\text{O}(\text{l}))] + [(1 \text{ mol})(\Delta H_f^\circ \text{C}_2\text{H}_2(\text{g})) + (\frac{13}{2} \text{ mol})(\Delta H_f^\circ \text{O}_2(\text{g}))]$ B. $\Delta H^\circ = [(1 \text{ mol})(\Delta H_f^\circ \text{C}_2\text{H}_2(\text{g})) + (\frac{13}{2} \text{ mol})(\Delta H_f^\circ \text{O}_2(\text{g}))] - [(4 \text{ mol})(\Delta H_f^\circ \text{CO}_2(\text{g})) + (5 \text{ mol})(\Delta H_f^\circ \text{H}_2\text{O}(\text{l}))]$

C. $\Delta H^\circ = [(1 \text{ mol})(\Delta H_f^\circ \text{C}_2\text{H}_2(\text{g})) + (\frac{13}{2} \text{ mol})(\Delta H_f^\circ \text{O}_2(\text{g}))] - [(4 \text{ mol})(\Delta H_f^\circ \text{CO}_2(\text{g})) + (5 \text{ mol})(\Delta H_f^\circ \text{H}_2\text{O}(\text{l}))]$

D. $\Delta H^\circ = [(4 \text{ mol})(\Delta H_f^\circ \text{CO}_2(\text{g})) + (5 \text{ mol})(\Delta H_f^\circ \text{H}_2\text{O}(\text{l}))] - [(1 \text{ mol})(\Delta H_f^\circ \text{C}_2\text{H}_2(\text{g})) + (\frac{13}{2} \text{ mol})(\Delta H_f^\circ \text{O}_2(\text{g}))]$

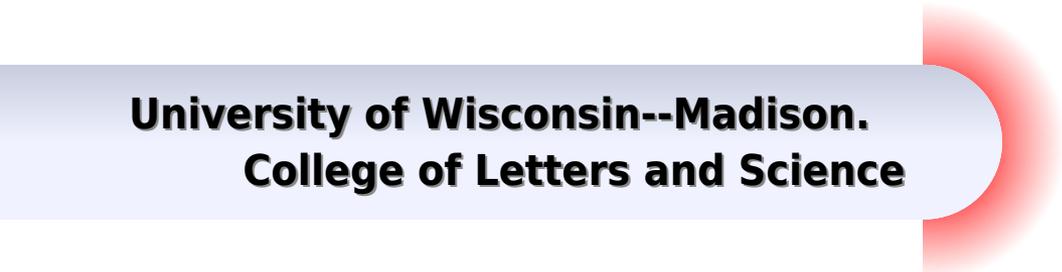
E. $\Delta H^\circ = [(1 \text{ mol})(\Delta H_f^\circ \text{CO}_2(\text{g})) + (\frac{4}{5} \text{ mol})(\Delta H_f^\circ \text{H}_2\text{O}(\text{l}))] - [(1 \text{ mol})(\Delta H_f^\circ \text{C}_2\text{H}_2(\text{g})) + (\frac{13}{2} \text{ mol})(\Delta H_f^\circ \text{O}_2(\text{g}))]$
- For which of the following statements is correct for the reaction

$$2\text{H}_2\text{O}_2(\text{g}) \rightarrow \text{H}_2(\text{g}) + 2\text{O}_2(\text{g}) + 66.4 \text{ kJ}$$

A. exothermic, potential energy greater in reactants
 B. exothermic, potential energy greater in products
 C. endothermic, potential energy greater in products
 D. endothermic, potential energy less in reactants
 E. endothermic, potential energy less in products

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The Top Books of the Year Chemistry For High School 10 12 The year 2023 has witnessed a remarkable surge in literary brilliance, with numerous compelling novels captivating the hearts of readers worldwide. Lets delve into the realm of bestselling books, exploring the engaging narratives that have enthralled audiences this year. The Must-Read : Colleen Hoover "It Ends with Us" This heartfelt tale of love, loss, and resilience has captivated readers with its raw and emotional exploration of domestic abuse. Hoover expertly weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can succeed. Chemistry For High School 10 12 : Taylor Jenkins Reids "The Seven Husbands of Evelyn Hugo" This captivating historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reids captivating storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery. Discover the Magic : Delia Owens "Where the Crawdads Sing" This evocative coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens crafts a tale of resilience, survival, and the transformative power of nature, captivating readers with its evocative prose and mesmerizing setting. These top-selling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of compelling stories waiting to be discovered. The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is a brilliant and gripping novel that will keep you guessing until the very end. The novel is a warning tale about the dangers of obsession and the power of evil.

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Table of Contents Chemistry For High School 10 12

1. Understanding the eBook Chemistry For High School 10 12
 - The Rise of Digital Reading Chemistry For High School 10 12
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry For High School 10 12
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry For High School 10 12
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry For High School 10 12
 - Personalized Recommendations
 - Chemistry For High School 10 12 User Reviews and Ratings
 - Chemistry For High School 10 12 and Bestseller Lists
5. Accessing Chemistry For High School 10 12 Free and Paid eBooks
 - Chemistry For High School 10 12 Public Domain eBooks
 - Chemistry For High School 10 12 eBook Subscription Services
 - Chemistry For High School 10 12 Budget-Friendly Options
6. Navigating Chemistry For High School 10 12 eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry For High School 10 12 Compatibility with Devices
 - Chemistry For High School 10 12 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry For High School 10 12
 - Highlighting and Note-Taking Chemistry For High School 10 12
 - Interactive Elements Chemistry For High School 10 12
8. Staying Engaged with Chemistry For High School 10 12

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry For High School 10 12
9. Balancing eBooks and Physical Books Chemistry For High School 10 12
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry For High School 10 12
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry For High School 10 12
- Setting Reading Goals Chemistry For High School 10 12
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry For High School 10 12
- Fact-Checking eBook Content of Chemistry For High School 10 12
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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