

Grade 11 Chemistry Exam Review

1. Fill in the following chart with the number of electrons, protons, neutrons, atomic number, mass number, and the element notations

Atom	Atomic Number	Mass Number	# of protons	# of neutrons	# of electrons
K ⁺					
Es					
Ms					
P ³⁻	13		13	14	
		75		42	
		184	74	117	78
	99	252			
		280		157	

2. Draw Bohr diagrams and Lewis dot diagrams for the following elements: Li, Ca, Al, C, P, O, F, K
3. Draw Bohr diagrams of the following ions: Na⁺, Mg²⁺, B³⁻, Sr²⁺, N³⁻, S²⁻, Cl⁻
4. Find the average atomic mass of P when the atomic mass with 15 neutrons is 30 is in 77% abundance and the atomic mass with 17 neutrons is 32 is in 23% abundance. (30.46 amu)
5. Find the Average atomic mass of Ca when the element which has 20 neutrons is in 12% abundance and the element which has 21 neutrons is in 88% abundance. (40.88 amu)
6. Find the average atomic mass of Cr if the atomic number is 24 and the two isotopes are; 18 neutrons at 97% abundance, and 19 neutrons at 3% abundance. (42.03 amu)
7. What is ionization energy, electron affinity, and electronegativity?
8. For the following elements decide which one has a larger atomic radius, electronegativity, ionization energy, and reactivity.

		Atomic Radius largest	electronegativity largest	ionization energy largest	reactivity greatest
O	Po				
B	N				
Ni	Pt				
Mg	Ra				
K	Fe				
Ag	Rb				
Au	Cu				
Hg	Zn				
I	Cl				
As	N				

9. Describe how the atomic radius gets smaller as we add more electrons in a row.
10. Describe how the reactivity is linked to ionization energy for the metals and electron affinity for the non-metals.
11. Which family of elements do the following elements belong?

a) He _____	b) Cl _____
c) Ca _____	d) Cs _____
e) Fr _____	f) Es _____
g) Yb _____	h) I _____
i) Np _____	j) Sr _____
k) Fr _____	l) Xe _____
12. What is the octet rule and what does it mean for any element that has its octet full.
13. What is the bonding capacity of the following elements: P, S, Cl, Na, K, Ba, B, Mg, Li, Si
14. Define the following terms: alloy, metalloid, solution, hydrate
15. How is the matter classified?
16. Describe the difference between ionic bonds and covalent bonds. Be sure to include which type of elements are joined together for each type of bond, how strong the bond is, if the electrons are either transferred or shared.
17. Indicate whether the following are ionic or covalent bonds and draw the lewis structures would look with the bonds.

a) LiF	b) F ₂	c) MgS	d) O ₂
e) AlF ₃	f) CO ₂	g) H ₂ O	h) Mg ₃ N ₂

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**Thirumalaisamy P.
Velavancorresponding**



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