

## CHAPTER 7 PRACTICE TEST: ATOMIC STRUCTURE &amp; PERIODIC TRENDS- KEY

1) When a solution of sodium chloride is vaporized in a flame, the color of the flame is:

- A) Blue      **B) Yellow**      C) Green      D) Violet      E) White

2) Electromagnetic radiation travels through vacuum at a speed of \_\_\_\_\_ m/s.

- A) 186,000      **C)  $3.00 \times 10^8$**       E) It depends on  $\lambda$   
 B) 125      D) 10,000

3) The wavelength of light that has a frequency of  $1.20 \times 10^{13} \text{ s}^{-1}$  is \_\_\_\_\_ m.

- A) 25.0      **B)  $2.50 \times 10^{-2}$**       C) 0.0400      E) 2.5  
 D) 12.0

GIVEN:  $\nu = 1.20 \times 10^{13} \text{ s}^{-1}$

UNKNOWN:  $\lambda$

CONNECTION:  $c = \lambda \nu$

SUBSTITUTE & SOLVE:  $\lambda = c/\nu$

$$\lambda = \frac{3.00 \times 10^8 \text{ m} \cdot \text{s}^{-1}}{1.20 \times 10^{13} \text{ s}^{-1}}$$

$$\lambda = 2.50 \times 10^{-2} \text{ m}$$

4) Ham radio operators often broadcast on the 6-meter band. The frequency of this EM radiation is \_\_\_\_\_ MHz.

- A) 500      B) 200      **C) 50**      D) 20      E) 2.0

GIVEN:  $\lambda = 6 \text{ m}$

UNKNOWN:  $\nu$  (Must convert answer to MHz!)

CONNECTION:  $c = \lambda \nu$

SUBSTITUTE & SOLVE:  $\nu = c/\lambda$

$$\nu = \frac{3.00 \times 10^8 \text{ m} \cdot \text{s}^{-1}}{6 \text{ m}}$$

$$\nu = 0.5 \times 10^8 \text{ Hz} \times \frac{1 \text{ MHz}}{10^6 \text{ Hz}} = 50 \text{ MHz}$$

5) The energy of a photon of light is \_\_\_\_\_ proportional to its frequency and \_\_\_\_\_ proportional to its wavelength.

- A) directly, directly      C) inversely, directly      E) indirectly, not  
 B) inversely, inversely      **D) directly, inversely**

6) The energy of a photon that has a wavelength of 12.3 nm is \_\_\_\_\_ J.

- A)  $1.51 \times 10^{-17}$       B)  $4.42 \times 10^{-23}$       C)  $1.99 \times 10^{-23}$       D)  $2.72 \times 10^{-20}$       **E)  $1.62 \times 10^{-17}$**

7) The wavelength of a photon that has an energy of  $5.25 \times 10^{-19} \text{ J}$  is \_\_\_\_\_ m.

- A)  $3.79 \times 10^{-7}$**       B)  $2.64 \times 10^6$       C)  $2.38 \times 10^{23}$       D)  $4.21 \times 10^{-24}$       E)  $3.79 \times 10^7$

8) A mole of red photons of wavelength 725 nm has \_\_\_\_\_ kJ of energy.

- A)  $2.74 \times 10^{-19}$       B)  $4.56 \times 10^{-26}$       C)  $6.05 \times 10^{-7}$       **D) 165**      E) 227

GIVENS:  $\lambda = 725 \text{ nm}$  (must convert to m before using formulas)

$$\lambda = 725 \text{ nm} \times \frac{1 \text{ m}}{10^9 \text{ nm}} = 7.25 \times 10^{-7} \text{ m}$$

# Holt Chemistry Chapter 7 Test

**Patrick Vollmar**



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**Holt Chemistry** Ralph Thomas Myers,2004 *Books in Print Supplement* ,2002 Holt Chemistry File ,1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket **Children's Books in Print, 2007** ,2006 Holt Biology Holt Rinehart & Winston,2003-08 *Children's Books in Print* R R Bowker Publishing,Bowker,1999-12 **El-Hi Textbooks and Serials in Print** ,1985 Holt Algebra 1 2003 Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2003 **El-Hi Textbooks and Serials in Print, 1996** R R Bowker Publishing,Bowker,1996-02 **Trace Environmental Quantitative Analysis** Paul R. Loconto,2020-12-27 A thorough and timely update this new edition presents principles techniques and applications in this sub discipline of analytical chemistry for quantifying traces of potentially toxic organic and inorganic chemical substances found in air soil fish and water as well as serum plasma urine and other body fluids The author addresses regulatory aspects calibration verification and the statistical treatment of analytical data including instrument detection limits quality assurance quality control sampling and sample preparation and techniques that are used to quantify trace concentrations of organic and inorganic chemical substances Key Features Fundamental principles are introduced for the more significant experimental approaches to sample preparation Principles of instrumental analysis determinative techniques for trace organics and trace inorganics analysis An introduction to the statistical treatment of trace analytical data How to calculate instrument detection limits based on weighted least squares confidence band calibration statistics Includes an updated series of student tested experiments

**Keys to ... Tests in Chemistry Series A and Series B** Charles Elwood Dull,1958 *El-Hi Textbooks & Serials in Print, 2005* ,2005 *Construction of a Diagnostic Test in the Mechanics and Related Fundamentals of High School Chemistry ...* Joseph Earl Malin,1932 **A Detailed Analysis of Achievement in the High School** Cecile White Flemming,1925

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