

# Soluzioni

## CAPITOLO 15

### VERIFICA LE TUE CONOSCENZE

#### VALENZA E NUMERO DI OSSIDAZIONE

- 1**
- La valenza di un elemento è il numero di elettroni che l'atomo cede, acquista o condivide quando si lega ad altri atomi.
  - La valenza massima corrisponde al numero romano che contraddistingue il gruppo, quindi, nel caso di un elemento del terzo gruppo della tavola periodica, è 3.
  - Un atomo di carbonio può formare al massimo 4 legami.
- 2** (C) di cariche reali o convenzionali associato a ciascun atomo del composto.
- 3** Regole per il calcolo del numero di ossidazione non corrette [tra parentesi, la versione corretta]
- l'ossigeno ha sempre n.o.  $-2$ ; [l'ossigeno nei composti ha n.o.  $-2$ , tranne nei perossidi, nei superossidi e nel difluoruro di ossigeno, dove ha rispettivamente n.o.  $-1$ ,  $-1/2$ ,  $+2$ ];
  - l'idrogeno ha n.o.  $+1$  nei composti e  $+2$  nella molecola  $H_2$ ; [l'idrogeno nei composti ha n.o.  $+1$  quando è legato a un non metallo e  $-1$  quando è legato a un metallo].

#### SCRIVERE LE FORMULE DEI COMPOSTI A PARTIRE DAI NUMERI DI OSSIDAZIONE

- 4**  $Na_2O$ : in questo modo la somma algebrica dei n.o. è uguale a zero:  $2 \cdot (+1) + (-2) = 0$ .
- 5**  $Ca_3(PO_4)_2$
- 6**
- $SO_4^{2-}$ ; carica dello ione =  $-2$
  - $ClO_7^-$ ; carica dello ione =  $-1$
- 7** La formula  $Cl_2Ba$  è sbagliata, in quanto si scrive a sinistra l'atomo con n.o. positivo ( $Ba_2Cl_2$ ).
- 8** È corretta soltanto: d.  $SO_2$  [tra parentesi, le formule corrette]
- HS [ $H_2S$ ]
  - $Fe_3S_2$  [ $Fe_2S_3$ ]
  - $LiS_2$  [ $Li_2S$ ]

#### LA NOMENCLATURA DEI COMPOSTI BINARI

- 9** Un composto binario è formato da due elementi diversi. I composti binari si dividono in:
- composti contenenti ossigeno: ossidi (es.  $ZnO$ ), perossidi (es.  $H_2O_2$ ), superossidi;
  - composti contenenti idrogeno: idruri (es.  $NaH$ ), idracidi (es.  $HF$ );
  - sali binari (es.  $NaCl$ ,  $CaCl_2$ ).
- 10** In molti composti binari si ha la desinenza *-ure*; tale desinenza non compare negli ossidi.
- 11** La notazione di Stock si utilizza quando gli elementi che formano un composto hanno più di un numero di ossidazione. Comporta l'eliminazione dei prefissi numerici e l'introduzione del numero di ossidazione dell'elemento, come numero romano tra parentesi posto dopo il nome dell'elemento. Si ritiene che in questo modo il nome che contraddistingue l'elemento risulti più breve e più chiaro.
- (n.o. Fe =  $+2$ )     $FeBr_2$     bromuro di ferro(II)  
(n.o. Fe =  $+3$ )     $FeBr_3$     bromuro di ferro(III)
- 12** Lo ione ossido ( $O^{2-}$ ) è formato da un atomo di O con n.o. =  $-2$ .  
Lo ione perossido ( $O_2^{2-}$ ) è formato da due atomi di O entrambi con n.o. =  $-1$ , legati da legame covalente.  
Lo ione superossido ( $O_2^-$ ) è formato da due atomi di O con n.o. medio =  $-1/2$ , legati da legame covalente.  
Un ossido acido è formato da un non metallo e ossigeno, un ossido basico è formato da un metallo e ossigeno.
- 13**
- Idrogeno e metallo: idruri ( $NaH$ ,  $LiH$ );
  - idrogeno e non metallo dei gruppi 16 e 17: idracidi ( $HCl$ ,  $HF$ ,  $HBr$ );
  - metallo e non metallo: sali binari ( $NaCl$ ,  $FeCl_2$ ).
- 14** (B) cianidrico.
- 15**  $K_2O$ : ossido di potassio;  
 $SO_2$ : anidride solforosa (n.o. S =  $+4$ ).  
 $K_2O$  si chiama *ossido* perché è un ossido basico (metallo e ossigeno).  
 $SO_2$  si chiama *anidride* perché è un ossido acido (non metallo e ossigeno).

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