

Atomic Structure Worksheet

PART A – SUBATOMIC PARTICLES

The table below contains information about several elements. In each case, enough information has been provided for you to fill in the blanks. Assume all atoms are neutral.

| Isotope Name | Nuclear Symbol | Atomic Number | Mass Number | # of Protons | # of Electrons | # of Neutrons |
|---------------|------------------------|---------------|-------------|--------------|----------------|---------------|
| 1. calcium-40 | | | | | | |
| 2. | | 12 | 24 | | | |
| 3. | | | | 1 | | 2 |
| 4. | $^{197}_{79}\text{Au}$ | | | | | |
| 5. | | | | | 26 | 30 |
| 6. | | | 201 | 80 | | |
| 7. | | 17 | | | | 18 |

PART B – AVERAGE ATOMIC MASS

- Calculate the average atomic mass for neon if its abundance in nature is 90.5% neon-20, 0.3% neon-21, and 9.2% neon-22.
- Calculate the average atomic mass of silver if 13 out of 25 atoms are silver-107 and 12 out of 25 atoms are silver-109.
- Distinguish between mass number, relative atomic mass, and average atomic mass.

Chapter 4 Atomic Structure Guided Practice Problems Answers

Guillaume Favre



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