

**CHAPTER 6 REVIEW***Chemical Bonding***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

\_\_\_\_\_ **lattice energy** \_\_\_\_\_ a. ionic bond  
 \_\_\_\_\_ **bond energy** \_\_\_\_\_ b. covalent bond  
 \_\_\_\_\_ **enthalpy of vaporization** \_\_\_\_\_ c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

\_\_\_\_\_ **ionic** \_\_\_\_\_ a. H—F \_\_\_\_\_ **nonpolar covalent** \_\_\_\_\_ d. H—H  
 \_\_\_\_\_ **ionic** \_\_\_\_\_ b. Na—Cl \_\_\_\_\_ **polar covalent** \_\_\_\_\_ e. H—C  
 \_\_\_\_\_ **polar covalent** \_\_\_\_\_ c. H—O \_\_\_\_\_ **polar covalent** \_\_\_\_\_ f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H<sub>2</sub>S and H<sub>2</sub>O have similar structures and their central atoms belong to the same group. Yet H<sub>2</sub>S is a gas at room temperature and H<sub>2</sub>O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increased polarity in H<sub>2</sub>O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

# Chapter 8 Covalent Bonding Study Answers

**Robinson**



## Chapter 8 Covalent Bonding Study Answers:

*Study Guide/Selected Solutions Manual* Julie R. Frentrop, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

*Investigating Chemistry* Matthew Johll, 2008-12-22 In its new second edition *Investigating Chemistry A Forensic Science Perspective* remains the only book that uses the inherently fascinating topics of crime and criminal investigations as a context for teaching the fundamental chemical concepts most often covered in an introductory nonmajors course Covering all the standard topics Matthew Johll capitalizes on the surge of interest in the scientific investigation of crime as sparked by CSI and other television shows bringing together the theme of forensic science and the fundamentals of chemistry in ways that are effective and accessible for students This edition features refined explanations of the chemical concepts which are the core of the book as well as a more thoroughly integrated forensic theme updated features and an expanded media supplements package

*Study Guide and Solutions Manual for Organic Chemistry Digital Update* K. Peter C. Vollhardt, Neil E. Schore, 2024-03-25 Written by Organic Chemistry coauthor Neil Schore this invaluable manual includes chapter introductions that highlight new materials chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

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**Study Guide with Solutions for Organic Chemistry, Third Edition** Ralph J. Fessenden,Joan S. Fessenden,1986 Provides a set of additional drill problems chapter by chapter discussions and supplemental instructional material to help students master organic chemistry problem solving techniques

*Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e* Kenneth W. Raymond,2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

Robinson Chemistry Study Guide Robinson,1992

**Problem Solving in General Chemistry** Kenneth W. Whitten,1984

**Organic Chemistry, Study Guide and Solutions Manual** T. W. Graham Solomons,Craig B. Fryhle,2000 On the cover of this book is a Pacific yew tree found in the ancient forests of the Pacific Northwest The bark of the Pacific yew tree produces Taxol found to be a highly effective drug against ovarian and breast cancer Taxol blocks mitosis during eukaryotic cell division The supply of Taxol from the Pacific yew tree is vanishingly small however A single 100 year old tree provides only about one dose of the drug roughly 300 mg For this reason as well as the spectacular molecular architecture of Taxol synthetic organic chemists fiercely undertook efforts to synthesize it Five total syntheses of Taxol have thus far been reported Now a combination of isolation of a related metabolite from European yew needles and synthesis of Taxol from that intermediate supply the clinical demand This case clearly demonstrates the importance of synthesis and the use of organic chemistry It s just one of the many examples used in the text that will spark the interest of students and get them involved in the study of organic chemistry

**Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today** Seager/Slabaugh,2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

*General Chemistry* James E. Brady,Gerard E. Humiston,1982 The Fifth Edition retains the pedagogical strengths that made the previous editions so popular and has been updated

reorganized and streamlined Changes include more accessible introductory chapters with greater stress on the logic of the periodic table earlier introduction of redox reactions greater emphasis on the concept of energy a new section on Lewis structures earlier introduction of the ideal gas law and a new development of thermodynamics Each chapter ends with review questions and problems

**Chemistry and Industry Review** ,1958 *Selected Solutions for Chemistry, Concepts and Models* by Robinson, Odom, and Holtzclaw John H. Meiser,1992 *Chemical Molecular Science* Conrad L. Stanitski,2004-08 *Fundamentals of Organic Chemistry, Textbook, Study Guide and Solutions Manual* T. W. Graham Solomons,1996-08-15 A realistic approach to the study of mechanisms The book addresses real functional group chemistry with an emphasis on the biological environmental and medical applications of organic chemistry **MCAT Comprehensive Review** ,2004

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