

**1. Zeichnen Sie die folgenden Komplexe, geben Sie die Oxidationszahl des Zentralatoms an!**

- a)  $[PtCl_4]^{2-}$  → Oxidationszahl Zentralatom: +2
- b)  $[PtCl_6]^{2-}$  → Oxidationszahl Zentralatom: +4
- c)  $[PtCl_2(NH_3)_4]^{2+}$  → Oxidationszahl Zentralatom: +4
- d)  $[PtCl_2(NH_3)_2]^{2+}$  → Oxidationszahl Zentralatom: +4
- e)  $[PtCl_2(NH_3)_2]^{2+}$  → Oxidationszahl Zentralatom: +4
- f)  $[PtCl_2(NH_3)_2]^{2+}$  → Oxidationszahl Zentralatom: +4
- g)  $[PtCl_2(NH_3)_2]^{2+}$  → Oxidationszahl Zentralatom: +4
- h)  $[PtCl_2(NH_3)_2]^{2+}$  → Oxidationszahl Zentralatom: +4

Zeichnen Sie die folgenden Komplexe, geben Sie die Oxidationszahl des Zentralatoms an!

- a)  $[PtCl_2(NH_3)_2]^{2+}$  → Oxidationszahl Zentralatom: +4
- b)  $[PtCl_2(NH_3)_2]^{2+}$  → Oxidationszahl Zentralatom: +4
- c)  $[PtCl_2(NH_3)_2]^{2+}$  → Oxidationszahl Zentralatom: +4

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- a)  $[PtCl_2(NH_3)_2]^{2+}$  → Oxidationszahl Zentralatom: +4
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- h)  $[PtCl_2(NH_3)_2]^{2+}$  → Oxidationszahl Zentralatom: +4

# Chapter 8 Solutions Acids Bases Assessment Answer Key

**Yijin Wang**



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