

CHEM 110 Exam 1 - Practice Test 1 Solutions

1D

The number of protons is always equal to the Atomic number which is the number found on the periodic table above the symbol. The mass number is the number of neutrons added to the number of protons ($\text{Mass} = \#p^+ + \#n^0$). In this problem the mass number is 80. We know that the number of protons is equal to 35, so the number of neutrons must be 45 ($80 = 35p^+ + \#n^0$). The charge on an atom is the number of protons minus the number of electrons. ($\text{Charge} = \#p^+ - \#e^-$). Here, we know the charge is -1 so $(-1 = 35 - \#e^-) = 36$. This makes sense because we would assume that there should be more electrons since the overall charge is negative.

2B

s orbitals are regular circles like A. p orbitals are 8 shaped like B. D orbitals are shaped like answer D. You will learn later in class about hybrid orbitals, which can look like answer C. Bromine on the periodic table is found in the 3-8 group. We know that these groups are filled with electron in the p orbitals. The last electron in the Bromine atom is part of a p orbital so the answer is B.

3C

Keep in mind that for elements other than hydrogen, the 4s is always filled before the 3d. The electron configuration shows the valence electrons to be in the 4 row/family of the periodic table. We see that the s orbital is filled and the d orbital has 8 electrons in it. The only choice is Ni. Answers A and D are tricks. They both have the same number of electrons but different configuration. Remember that when ionizing an atom other than Hydrogen, always take from the highest energy level, not the last filled.

4D

The l quantum number refers to the s orbital. Fe^{3+} has 0 valence electrons in the s orbital, but it has 6 in the core electrons. Keep in mind that electrons in the 4s are ionized before 3d.

5C

This question is based off a periodic trend. Reactivity increases as you go down a group. Therefore this answer is based on simply going down the correct group. Group 1 is the left most group.

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Rex M. Heyworth



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