

AP Chemistry: Chapter 14 Practice Test

10. Given the following acids and K_a values:

HClO_4	HOAc	HCN	HF
-----------------	---------------	--------------	-------------

1×10^7	1.76×10^{-5}	4.93×10^{-10}	3.53×10^{-4}
-----------------	-----------------------	------------------------	-----------------------

which shows the conjugate bases listed by increasing strength?

- | | |
|--|--|
| [A] ClO_4^- , F^- , OAc^- , CN^- | [B] CN^- , F^- , OAc^- , ClO_4^- |
| [C] CN^- , OAc^- , F^- , ClO_4^- | [D] CN^- , ClO_4^- , F^- , OAc^- |
| [E] ClO_4^- , OAc^- , CN^- , F^- | |
15. Given that the K_a for HOCl is 3.5×10^{-8} , calculate the K value for the reaction of HOCl with OH^- .
- | | | |
|---------------------------|-----------------------|--------------------------|
| [A] 3.5×10^{-22} | [B] 3.5×10^6 | [C] 2.9×10^{-7} |
| [D] 3.5×10^{-8} | [E] none of these | |
32. Acetic acid ($\text{HC}_2\text{H}_3\text{O}_2$) is a weak acid ($K_a = 1.8 \times 10^{-5}$). Calculate the pH of a 17.6 M $\text{HC}_2\text{H}_3\text{O}_2$ solution.
- | | | | | |
|---------|---------|---------|----------|---------|
| [A] 4.3 | [B] 6.4 | [C] 1.7 | [D] 0.97 | [E] 7.4 |
|---------|---------|---------|----------|---------|
52. What concentration of acetic acid ($K_a = 1.80 \times 10^{-5}$) has the same pH as that of 5.00×10^{-3} M HCl ?
- | | | |
|-----------------------------|-----------------------------|------------|
| [A] 5.00×10^{-3} M | [B] 3.60×10^{-3} M | [C] 1.39 M |
| [D] 5.00 M | [E] none of these | |

Consider the reaction $\text{HOCl} + \text{F}^- \rightleftharpoons \text{HF} + \text{OCl}^-$

Chemistry Chapter 14 Test

Daniel C. Harris



Chemistry Chapter 14 Test:

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Table of Contents Chemistry Chapter 14 Test

1. Understanding the eBook Chemistry Chapter 14 Test
 - The Rise of Digital Reading Chemistry Chapter 14 Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 14 Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 14 Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 14 Test
 - Personalized Recommendations
 - Chemistry Chapter 14 Test User Reviews and Ratings
 - Chemistry Chapter 14 Test and Bestseller Lists

5. Accessing Chemistry Chapter 14 Test Free and Paid eBooks
 - Chemistry Chapter 14 Test Public Domain eBooks
 - Chemistry Chapter 14 Test eBook Subscription Services
 - Chemistry Chapter 14 Test Budget-Friendly Options
6. Navigating Chemistry Chapter 14 Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 14 Test Compatibility with Devices
 - Chemistry Chapter 14 Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 14 Test
 - Highlighting and Note-Taking Chemistry Chapter 14 Test
 - Interactive Elements Chemistry Chapter 14 Test
8. Staying Engaged with Chemistry Chapter 14 Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 14 Test
9. Balancing eBooks and Physical Books Chemistry Chapter 14 Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 14 Test
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 14 Test
 - Setting Reading Goals Chemistry Chapter 14 Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 14 Test
 - Fact-Checking eBook Content of Chemistry Chapter 14 Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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