

**Chemical Equilibrium****STUDY QUESTIONS & PROBLEMS**

1. Write the expressions for the equilibrium constant
- $K_c$
- for the following reactions:



2. Write the equilibrium constant expressions for the following reactions. How are they related to one another?



3. Calculate the value of the equilibrium constant for the following system, given the data shown:



Concentrations at equilibrium:

$[\text{H}_2]$	=	1.5 mol Liter <sup>-1</sup>
$[\text{CO}_2]$	=	2.5 mol Liter <sup>-1</sup>
$[\text{H}_2\text{O}]$	=	0.5 mol Liter <sup>-1</sup>
$[\text{CO}]$	=	3.0 mol Liter <sup>-1</sup>

4. Chlorine molecules will dissociate at high temperatures into chlorine atoms. At 3000°C, for example,  $K_c$  for the equilibrium shown is 0.55. If the partial pressure of chlorine molecules is 1.5 atm, calculate the partial pressure of the chlorine atoms:  

$$\text{Cl}_2(\text{g}) \rightleftharpoons 2 \text{Cl}(\text{g})$$
5. Suppose that 0.50 moles of hydrogen gas, 0.50 moles of iodine gas, and 0.75 moles of hydrogen iodide gas are introduced into a 2.0 Liter vessel and the system is allowed to reach equilibrium.  

$$\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2 \text{HI}(\text{g})$$
  
 Calculate the concentrations of all three substances at equilibrium. At the temperature of the experiment,  $K_c$  equals  $2.0 \times 10^{-2}$ .
6. If the mechanism of a chemical equilibrium consists of two reversible elementary steps, each with its own equilibrium constant  $K_{c1}$  and  $K_{c2}$ , what expression relates the equilibrium constant  $K_c$  for the overall equilibrium to the two constants  $K_{c1}$  and  $K_{c2}$ ?
7. When 2.0 mol of carbon disulfide and 4.0 mol of chlorine are placed in a 1.0 Liter flask, the following equilibrium system results. At equilibrium, the flask is found to contain 0.30 mol of carbon tetrachloride. What quantities of the other components are present in this equilibrium mixture?



# Chemical Equilibrium Practice Problems And Solutions

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## **Chemical Equilibrium Practice Problems And Solutions:**

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*Student Solutions Manual* Darrell Ebbing, 2005-03-17 Provides solutions to odd numbered Practice Problems General Problems and Cumulative Skills Problems plus answers to Review Questions *Solutions Manual for Quanta, Matter and Change* Peter Atkins, Julio dePaula, Ron Friedman, 2008-12-15 **Chemical Engineering Principles and Applications** Nuggenhalli S. Nandagopal, 2023-05-26 This text provides a clear and concise understanding of the principles and applications of chemical engineering using a rigorous yet easy to follow presentation The coverage is broad and it includes all the relevant concepts such as mass and energy balances mass transfer chemical reaction engineering and many more Elucidation of the principles is further reinforced by examples and practice problems with detailed solutions Firmly grounded in the fundamentals the book maximizes readers capacity to take on new problems and challenges in the field with confidence and conviction Providing a ready reference and review of essential principles and their applications in chemical engineering the book is ideal for undergraduate chemical engineering students as well as practicing engineers preparing for the engineering license exams FE and PE in USA and abroad *Instructor's Solutions Manual to Accompany Atkins' Physical Chemistry, Ninth Edition* C. A. Trapp, M. P. Cady, Carmen Giunta, 2010 The Instructor's solutions manual to accompany Atkins Physical Chemistry provides detailed solutions to the b exercises and the even numbered discussion questions and problems that feature in the ninth edition of Atkins Physical Chemistry The manual is intended for instructors and consists of material that is not available to undergraduates The manual is free to all adopters of the main text *Student's Solutions Manual to Accompany Atkins' Physical Chemistry* C. A. Trapp, M. P. Cady, Carmen Giunta, 2010 This solutions manual provides the authors detailed solutions to exercises and problems in physical chemistry It comprises solutions to exercises at the end of each chapter and solutions to numerical theoretical and additional problems *GO TO Objective NEET 2021 Chemistry Guide 8th Edition* Disha Experts, *Partial Solutions Manual* Darrell D. Ebbing, David Bookin, 1998 **Questions and Answers about Block Scheduling** Donald D. Gainey, John M. Brucato, 1999 First Published in 1999 Routledge is an imprint of Taylor Francis an informa company *Study Questions and Problems in Inorganic Chemistry* Alexander Silverman, 1924 *Ebbing* Darrell D. Ebbing, 1993 **Student's Guide to Chemistry; a Modern Introduction** David Brooks, 1974 *Annual of the University of Deseret* University of Utah, 1901 *Nature* Sir Norman Lockyer, 1905 **Brass World** ,1922 **Mining and Chemical Engineering Review** ,1920 *Chemical Addresses Delivered at the Second Decennial Celebration of Clark University* Clark University (Worcester, Mass.), 1911 **Mining Science** ,1911 *Mining American* ,1911 **Cracking the GRE Chemistry Test** ,2002 The Princeton Review realizes that acing the GRE Chemistry Test is very different from getting straight A s in school We don t try to teach you everything there is to know about chemistry only the techniques you ll need to score higher on the test There s a big difference In

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web jun 1 2020 lorry crane since september 2013 all lorry crane operators have to be certified to operate the crane that is attached on the lorry the training equipped operators with relevant skills and knowledge to operate the lorry crane safely this is like the driving license that all drivers must have before they are allowed to drive on

**directory of accredited crane operator certification programs** - Nov 25 2022

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