

Equilibrium Constant - Worksheet 1_ANSWERS

1. Write the equilibrium expression for the following reactions:



$$K = \frac{[U]^u [V]^v}{[W]^w [X]^x}$$



$$K = \frac{[H_2][CO_2]}{[H_2O][CO]}$$



$$K = \frac{[Cl_2][CO]}{[COCl_2]}$$



$$K = \frac{[HCl]^2}{[H_2][Cl_2]}$$



$$K = \frac{[NO][CO_2]}{[NO_2][CO]}$$



$$K = \frac{[Zn^{2+}]}{[Ag^+]^2}$$

*Note: solids are not included



$$K = \frac{[H_2][C_2H_4]}{[C_2H_6]}$$

2. Consider the following equilibrium reaction:



The K_{eq} for this reaction at 25°C is 1.02. At equilibrium the concentration of HBr is 0.50 mol/L. Assuming H_2 and Br_2 are present in equal amounts, calculate the concentration of H_2 at equilibrium.

$$[H_2] = 0.50 \text{ mol/L}$$

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Yi-Tong Ma



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Report of the Annual Meeting of the South African Association for the Advancement of Science, 2005

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The Science Teacher, 2008. SCC Library has 1964 cur. **Illinois Chemistry Teacher**, 1992. **Current Index to Journals in Education, Semi-Annual Cumulation, January-June, 1977**. Educational Resources Information Center Staff, 1977-09. **Biology**, 1999.

Chemical Equilibrium William Guenther, 1975-12. The present work is designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry, biochemistry, geochemistry, and interdisciplinary environmental fields. The pedagogical strategy I have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature. The procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution assumptions. There is repeated

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