

**3 • Reactions in Aqueous Solution****STUDY QUESTIONS AND PROBLEMS**

1. Classify each of the following solutes as a strong electrolyte, weak electrolyte, or nonelectrolyte:

sugar	sodium hydroxide
common salt (NaCl)	hydrochloric acid
alcohol	copper sulfate
acetic acid	carbonic acid

2. Predict the solubility of the following salts:

- sodium sulfate
- potassium chromate
- silver bromide
- nickel (II) hydroxide
- aluminum nitrate
- barium sulfide
- ammonium acetate
- strontium iodide

3. Write the ions that are produced when the following substances dissolve in water:

- $Mg(OH)_2$
- $K_2SO_4$
- $NaHCO_3$
- $(NH_4)_3PO_4$
- $NaClO$

4. Predict whether or not the following reactions will lead to a precipitate. Write complete and net ionic equations for all the reactions.

- a. potassium chromate and lead acetate
- b. silver perchlorate and ammonium chloride
- c. potassium carbonate and copper acetate
- d. sodium fluoride and magnesium iodide
- e. barium nitrate and potassium sulfate

# Reactions In Aqueous Solution Worksheet Answers

**CO Houle**



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