

chemistry stoichiometry questions and answers

University of Binawan

1. Complete combustion of ammonia gas requires oxygen gas with equality reaction :



With Assuming that the oxygen level in the air is 20%, then the volume of air needed to burn 4 liters of ammonia gas is...

- A. 5 liters
- B. 7 liters
- C. 10 liters
- D. 35 liters

Answer: D

Discussion

Equality reaction equivalent :



Volume gas oxygen = $(7/4) \times 4$ liters = 7 liters

Volume oxygen = 20% volume air

7 liters = 0.2 volume of air

volume udara = $7/0.2 = 35$ litres

2. Every 1 liter of nitrogen gas is appropriate finished react with 2.5 liters of oxygen gas to form 1 liter of nitrogen oxide gas . If the volume is measured at the same temperature and pressure , then formula molecule the nitrogen oxide is

- A. N_2O
- B. NO
- C. N_2O_3 _ _ _
- D. N_2O_5 _ _

Answer: D

Discussion

Equality reaction :



The number of N atoms before reaction = Number of N atoms after reaction

$$4 = 2x$$

$$x = 2$$

Number of O atoms before reaction = Number of O atoms after reaction

$$10 = 2y$$

$$y = 5$$

So, formula molecule The nitrogen oxide is N_2O_5 .

3. Propane gas (C_3H_8) is burned perfect with reaction :



For burn perfect every liters of propane gas, oxygen gas is required as much

- A. 1.0 liters
- B. 3.0 liters

Stoichiometry Questions Answers

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Table of Contents Stoichiometry Questions Answers

1. Understanding the eBook Stoichiometry Questions Answers
 - The Rise of Digital Reading Stoichiometry Questions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Questions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Questions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Questions Answers
 - Personalized Recommendations

- Stoichiometry Questions Answers User Reviews and Ratings
- Stoichiometry Questions Answers and Bestseller Lists
- 5. Accessing Stoichiometry Questions Answers Free and Paid eBooks
 - Stoichiometry Questions Answers Public Domain eBooks
 - Stoichiometry Questions Answers eBook Subscription Services
 - Stoichiometry Questions Answers Budget-Friendly Options
- 6. Navigating Stoichiometry Questions Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry Questions Answers Compatibility with Devices
 - Stoichiometry Questions Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Questions Answers
 - Highlighting and Note-Taking Stoichiometry Questions Answers
 - Interactive Elements Stoichiometry Questions Answers
- 8. Staying Engaged with Stoichiometry Questions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Questions Answers
- 9. Balancing eBooks and Physical Books Stoichiometry Questions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Questions Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Stoichiometry Questions Answers
 - Setting Reading Goals Stoichiometry Questions Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Stoichiometry Questions Answers
 - Fact-Checking eBook Content of Stoichiometry Questions Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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