

## Titration of an Antacid

### Purpose:

In this lab you will determine the acid neutralizing power of a commercially available antacid using a technique called back-titration.

### Overview:

Hydrochloric acid is the acid secreted by the parietal cells in the stomach. The flow of this acid increases when food enters the stomach. If you eat too much your stomach may produce too much of the acid. This produces the condition called "heartburn" or indigestion. To relieve this condition, antacids can be swallowed to neutralize the excess acid. The key ingredient in any antacid is a base. This base reacts with the hydrochloric acid to neutralize it. The table below contains a list of some antacids and the bases that they contain.

Brand	Base
Pepto-Bismol	$\text{BiO}(\text{HOC}_6\text{H}_4\text{COO})$
Milk of Magnesia	$\text{Mg}(\text{OH})_2$
Rolaids	$\text{NaAl}(\text{OH})_2\text{CO}_3$ or $\text{CaCO}_3$
Tums	$\text{CaCO}_3$
Alka-Seltzer II	$\text{NaHCO}_3$ and $\text{KHCO}_3$
Maalox	$\text{Mg}(\text{OH})_2$ and $\text{Al}(\text{OH})_3$
Gaviscon	$\text{Al}(\text{OH})_3$

Acids are neutralized by these bases as illustrated below:



In this lab you will find the neutralizing capacity for a commercial antacid. To do this, you will first dissolve an antacid in a measured excess of hydrochloric acid. This acid will react with the antacid, but leave some unreacted hydrochloric acid. You will then

# Antacid Titration Lab Report Answers

**L Towne**



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