

Name _____ Date: _____ Mode: _____

Determining Molecular Formulas (True Formulas)

1. The empirical formula of a compound is NO_2 . Its molecular mass is 92 g/mol. What is its molecular formula?

$$\begin{array}{r} \text{NO}_2 \\ 16.00\text{g} \\ 16.00\text{g} \\ + 14.00\text{g} \\ \hline 46.00\text{g} \end{array} \quad \frac{92\text{g}}{46.0\text{g}} = 2 \quad \text{N}_2\text{O}_4$$

2. The empirical formula of a compound is CH_2 . Its molecular mass is 70 g/mol. What is its molecular formula?

$$\begin{array}{r} \text{CH}_2 \\ 12.01\text{g} \\ 1.01\text{g} \\ + 1.01\text{g} \\ \hline 14.03\text{g} \end{array} \quad \frac{70\text{g}}{14.03\text{g}} = 4.98 = 5 \quad \text{C}_5\text{H}_{10}$$

3. A compound is found to be 40.0% carbon, 6.7% hydrogen and 53.3% oxygen. Its molecular mass is 60. g/mol. What is its molecular formula?

$$\begin{array}{l} 40.0\% \text{ C} \cdot \frac{1\text{mol C}}{12.01\text{g C}} = 3.33\text{mol C} \frac{3.33}{3.33} = 1 \\ 6.7\% \text{ H} \cdot \frac{1\text{mol H}}{1.01\text{g H}} = 6.63\text{mol H} \frac{6.63}{3.33} = 2 \\ 53.3\% \text{ O} \cdot \frac{1\text{mol O}}{16.00\text{g O}} = 3.34\text{mol O} \frac{3.34}{3.33} = 1 \end{array} \quad \begin{array}{l} \text{CH}_2\text{O} = 50.02\text{g} \\ \frac{60\text{g}}{30\text{g}} = 2 \\ \text{C}_2\text{H}_4\text{O}_2 \end{array}$$

4. A compound is 64.9% carbon, 13.5% hydrogen and 21.6% oxygen. Its molecular mass is 74 g/mol. What is its molecular formula?

$$\begin{array}{l} 64.9\% \text{ C} \cdot \frac{1\text{mol C}}{12.01\text{g C}} = 5.40\text{mol C} \frac{5.40}{1.35} = 4 \\ 21.6\% \text{ O} \cdot \frac{1\text{mol O}}{16.00\text{g O}} = 1.35\text{mol O} \frac{1.35}{1.35} = 1 \\ 13.5\% \text{ H} \cdot \frac{1\text{mol H}}{1.01\text{g H}} = 13.4\text{mol H} \frac{13.4}{1.35} = 10 \end{array} \quad \begin{array}{l} \text{C}_4\text{H}_{10}\text{O} \\ \uparrow \\ \text{empirical} \\ \uparrow \\ \text{molecular} \end{array}$$

5. A compound is 54.5% carbon, 9.1% hydrogen and 36.4% oxygen. Its molecular mass is 88 g/mol. What is its molecular formula?

$$\begin{array}{l} 54.5\% \text{ C} \cdot \frac{1\text{mol C}}{12.01\text{g C}} = 4.53\text{mol C} \frac{4.53}{2.28} = 2 \\ 9.1\% \text{ H} \cdot \frac{1\text{mol H}}{1.01\text{g H}} = 9.00\text{mol H} \frac{9.00}{2.28} = 4 \\ 36.4\% \text{ O} \cdot \frac{1\text{mol O}}{16.00\text{g O}} = 2.28\text{mol O} \frac{2.28}{2.28} = 1 \end{array} \quad \begin{array}{l} \text{C}_2\text{H}_4\text{O} = 44 \\ \frac{88}{2} = 44 \\ \text{C}_4\text{H}_8\text{O}_2 \end{array}$$

Molecular Formula Worksheet Answer Key

Thomas Griffiths



Molecular Formula Worksheet Answer Key:

Molecular Formula Worksheet Answer Key Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Molecular Formula Worksheet Answer Key**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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