

Chapter 17: Chemical Equilibrium

Worksheet: K_c and K_p

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Section 1

1. Write the equilibrium constant expression for each reaction.

- 1. $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$
- 2. $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$
- 3. $\text{H}_2(g) + \text{I}_2(g) \rightleftharpoons 2\text{HI}(g)$
- 4. $\text{PCl}_5(g) \rightleftharpoons \text{PCl}_3(g) + \text{Cl}_2(g)$
- 5. $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$
- 6. $\text{CO}_2(g) + \text{H}_2(g) \rightleftharpoons \text{CO}(g) + \text{H}_2\text{O}(g)$
- 7. $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$
- 8. $\text{H}_2(g) + \text{Cl}_2(g) \rightleftharpoons 2\text{HCl}(g)$
- 9. $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$
- 10. $\text{CO}_2(g) + \text{H}_2(g) \rightleftharpoons \text{CO}(g) + \text{H}_2\text{O}(g)$
- 11. $\text{H}_2(g) + \text{I}_2(g) \rightleftharpoons 2\text{HI}(g)$
- 12. $\text{PCl}_5(g) \rightleftharpoons \text{PCl}_3(g) + \text{Cl}_2(g)$
- 13. $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$
- 14. $\text{CO}_2(g) + \text{H}_2(g) \rightleftharpoons \text{CO}(g) + \text{H}_2\text{O}(g)$
- 15. $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$
- 16. $\text{H}_2(g) + \text{Cl}_2(g) \rightleftharpoons 2\text{HCl}(g)$

Section 2

- 17. Calculate K_c for the reaction: $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$. At equilibrium, [N₂] = 0.10 M, [H₂] = 0.30 M, and [NH₃] = 0.20 M.
- 18. Calculate K_c for the reaction: $\text{H}_2(g) + \text{I}_2(g) \rightleftharpoons 2\text{HI}(g)$. At equilibrium, [H₂] = 0.10 M, [I₂] = 0.10 M, and [HI] = 0.20 M.
- 19. Calculate K_c for the reaction: $\text{PCl}_5(g) \rightleftharpoons \text{PCl}_3(g) + \text{Cl}_2(g)$. At equilibrium, [PCl₅] = 0.10 M, [PCl₃] = 0.20 M, and [Cl₂] = 0.20 M.
- 20. Calculate K_c for the reaction: $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$. At equilibrium, [SO₂] = 0.10 M, [O₂] = 0.10 M, and [SO₃] = 0.20 M.
- 21. Calculate K_c for the reaction: $\text{CO}_2(g) + \text{H}_2(g) \rightleftharpoons \text{CO}(g) + \text{H}_2\text{O}(g)$. At equilibrium, [CO₂] = 0.10 M, [H₂] = 0.10 M, [CO] = 0.10 M, and [H₂O] = 0.10 M.
- 22. Calculate K_c for the reaction: $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$. At equilibrium, [N₂] = 0.10 M, [O₂] = 0.10 M, and [NO] = 0.20 M.
- 23. Calculate K_c for the reaction: $\text{H}_2(g) + \text{Cl}_2(g) \rightleftharpoons 2\text{HCl}(g)$. At equilibrium, [H₂] = 0.10 M, [Cl₂] = 0.10 M, and [HCl] = 0.20 M.
- 24. Calculate K_c for the reaction: $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$. At equilibrium, [N₂] = 0.10 M, [H₂] = 0.30 M, and [NH₃] = 0.20 M.
- 25. Calculate K_c for the reaction: $\text{CO}_2(g) + \text{H}_2(g) \rightleftharpoons \text{CO}(g) + \text{H}_2\text{O}(g)$. At equilibrium, [CO₂] = 0.10 M, [H₂] = 0.10 M, [CO] = 0.10 M, and [H₂O] = 0.10 M.

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