

Quiz & Worksheet - Colligative Properties and Raoult's Law

1. What happens to vapor pressure when you add a solute to a solution?

- It lowers the vapor pressure.
- It has no effect.
- It raises the vapor pressure.
- It causes the reaction to reach equilibrium.

2. What is osmotic pressure?

- It is the minimum pressure that must be applied to a solution to stop osmosis from happening.
- It is the maximum pressure that must be applied to a solution to stop osmosis from happening.
- It is the maximum pressure of the vapor over a liquid at equilibrium.
- It is the minimum pressure of the vapor over a liquid at equilibrium.

3.

9.0g of urea (a nonelectrolyte) is dissolved in 360 g water to create a solution. The temperature of the solution is 25.0°C. If the vapor pressure of water at 25.0°C is 23.80 mmHg, calculate the vapor pressure of the urea solution at 25.0°C in mmHg.

The molar mass of urea 60g/mol and water is 18g/mol.

- 0.59
- 23.91
- 23.61
- 3.57

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Colligative Properties Of Solutions Worksheet Answers

Daniela Niemeyer



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In a global inundated with monitors and the cacophony of instant transmission, the profound energy and emotional resonance of verbal art usually disappear in to obscurity, eclipsed by the continuous barrage of sound and distractions. Yet, situated within the musical pages of **Colligative Properties Of Solutions Worksheet Answers**, a captivating perform of literary splendor that pulses with organic thoughts, lies an unforgettable trip waiting to be embarked upon. Published by a virtuoso wordsmith, that exciting opus manuals readers on an emotional odyssey, gently exposing the latent potential and profound influence embedded within the complex web of language. Within the heart-wrenching expanse with this evocative examination, we shall embark upon an introspective exploration of the book is key styles, dissect its fascinating publishing fashion, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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