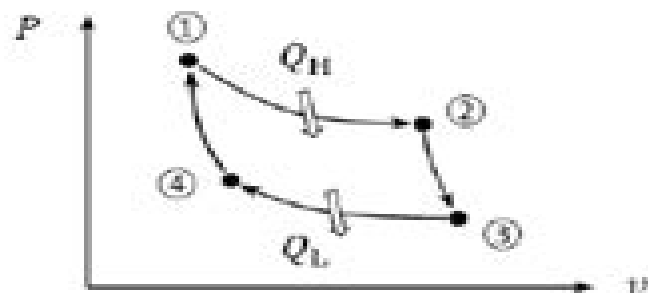


## Written Problem

The Carnot cycle is an idealization that involves four processes as shown (see also Figure 6–37 in textbook). As each of the processes is reversible, the entire cycle is also reversible.



- ① → ② : Isothermal expansion at  $T = T_H$
- ② → ③ : Adiabatic expansion (no heat transfer)
- ③ → ④ : Isothermal compression at  $T = T_L$
- ④ → ① : Adiabatic compression (no heat transfer)

- (a) Suppose the Carnot cycle is carried out in a closed system with an ideal gas as working fluid. Show that

$$Q_H = mRT_H \ln(V_2/V_1)$$

where  $m$  is the mass of the ideal gas. (Recall that the internal energy of an ideal gas depends *only* on the temperature  $T$ .)

You are now provided with more specifics: The ideal gas is 2 g of air. The Carnot cycle operates between  $T_H = 500$  K and  $T_L = 300$  K, with  $V_2 = 2V_1$ .

- (b) What is the amount of heat  $Q_L$  (kJ) that is transferred to the cold reservoir in every cycle?
- (c) Determine  $W_{\text{net}}$  (kJ), the net amount of work that is produced per cycle.

Note that  $W_{\text{net}}$  can be obtained by summing the  $\int P \, dV$  contribution from each of the four processes. This approach, however, is not recommended as there are much easier ways to get to the answer.

# Carnot Cycle Numerical Problems With Solutions

**Brendan G. Carr**



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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Carnot Cycle Numerical Problems With Solutions**," an enthralling opus penned by a highly acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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