



# COVALENT BONDING

## Practice Problems

Do your answers to the following problems.

### SECTION 5.1 MOLECULAR COMPOUNDS

- Classify each of the following as either an ionic or a molecular compound.
  - $\text{NaCl}$
  - $\text{H}_2\text{O}$
  - $\text{CaF}_2$
  - $\text{KNO}_3$
  - $\text{C}_2\text{H}_6$
  - $\text{H}_2\text{SO}_4$
  - $\text{MgO}$
  - $\text{H}_2\text{S}$
  - $\text{CO}_2$

- Classify each of the following as either an ionic or a molecular compound. Write the name of the compound.
- Write the formula for each of the following molecular compounds. Write the name of the compound.
- Write the formula for each of the following molecular compounds. Write the name of the compound.

### SECTION 5.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride.  $\text{H}-\text{F}$
- Show covalent bonding in methane.  $\text{C}$  with four  $\text{H}$  atoms bonded to it.
- Show covalent bonding in carbon dioxide.  $\text{C}$  with two  $\text{O}$  atoms bonded to it.
- Show covalent bonding in acetylene.  $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for each of the following. Show all covalent bonding explicitly.
  - $\text{H}_2\text{O}$
  - $\text{CO}_2$
  - $\text{CH}_4$
  - $\text{C}_2\text{H}_2$
  - $\text{C}_2\text{H}_4$
  - $\text{C}_2\text{H}_6$
  - $\text{C}_2\text{H}_2\text{O}_2$
  - $\text{C}_2\text{H}_4\text{O}_2$
  - $\text{C}_2\text{H}_6\text{O}_2$



### SECTION 5.3 BONDING: THE ORBITALS

- Which orbital and suborbital are represented by each of the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $3d$
  - $4s$
  - $4p$
  - $4d$
  - $4f$
- Which orbital and suborbital are represented by each of the following?
  - $1s$
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $3d$
  - $4s$
  - $4p$
  - $4d$
  - $4f$

# Chapter 8 Covalent Bonding Practice Problems Answers

**A Gutmann**



## **Chapter 8 Covalent Bonding Practice Problems Answers:**

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