

## MODULE C: SOLIDS, LIQUIDS & GASES

### Kinetic Molecular Theory & Phases of Matter

- The different phases of matter can be described based on the Kinetic Molecular Theory

a) According to Kinetic theory, an increase in temperature will cause the average kinetic energy of the particles in the solid, liquid, and gas to increase.

↑ temperature    ↑ average kinetic energy

b) When a solid is heated, the temperature is raised and the solid changes to a liquid at its melting point. If the heating is continued, the liquid will change into a gas at its boiling point.

c) The heat energy is changed into the equivalent amount of kinetic energy. Thus, the kinetic energy content of the particles increases in the order:

solid < liquid < gas

### - Kinetic Molecular Theory

- 1) Matter is made up of particles
- 2) The particles are always moving

### - Kinetic energy

→ energy of movement

### - Kinetic Molecular Theory for Gases

a) A gas consists of small particles (atoms or molecules) which have negligible volume ( $V$  of gas molecule  $\approx 0$ ). This implies that:

$$V_{\text{container}} = V_{\text{molecules}} + V_{\text{free space}} \approx V_{\text{free space}}$$

and that there is plenty of space for compression.

b) Gas particles are separated relatively far from one another.

c) Gas particles behave independently from one another. Thus, inter-particle forces of attraction are negligible.

d) The gas particles are in rapid and continuous motion.

e) In the gaseous state, the gas particles have vibrational, rotational, and translational motion.

1) Vibrational motion - a particle vibrates about a fixed point

2) Rotational motion - a particle spins about a fixed point

3) Translational motion - a particle moves from one place to another

f) Pressure of a gas comes from molecules colliding with the walls of the container.

Such collisions are perfectly elastic.

g) Average kinetic energy depends only on the absolute temperature in Kelvin (K)

$$\text{Average Kinetic Energy} = \frac{3}{2} RT, \quad R \rightarrow \text{gas law constant } (0.08206 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})$$

$T \rightarrow$  temperature (in Kelvin)

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**Puja Mehta**



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