

NAME \_\_\_\_\_ DATE \_\_\_\_\_ CLASS \_\_\_\_\_

**CHAPTER 12** **STUDY GUIDE FOR CONTENT MASTERY**

**Section 12.3 Limiting Reactants**

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of  $O_2$ ,  $NO$ , and  $NO_2$  are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of $O_2$	Amount of $NO$	Amount of $NO_2$	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	4 molecules	4 molecules	$NO$	2 molecules $O_2$
2 molecules	4 molecules	4 molecules	$O_2$	4 molecules $NO$
1.00 mol	2.00 mol	2.00 mol	none	none
4.00 mol	6.00 mol	4.00 mol	$NO$	2.00 mol $O_2$
0.05 mol	0.10 mol	0.10 mol	$NO$	1.50 mol $O_2$
1.00 mol	2.00 mol	2.00 mol	$O_2$	2.00 mol $NO$
0.200 mol	0.200 mol	0.400 mol	$NO$	0.400 mol $O_2$
32.00 g	60.02 g	92.02 g	none	none
64.00 g	60.02 g	66.02 g	$O_2$	30.12 g $NO$
64.00 g	30.01 g	38.76 g	$O_2$	1.24 g $NO$

balanced equation mole ratio = 2 mol  $NO$ /1 mol  $O_2$

16.00 g  $O_2$  = 1 mol  $O_2$ /32.00 g  $O_2$  = 0.5125 mol  $O_2$

20.00 g  $NO$  = 1 mol  $NO$ /30.01 g  $NO$  = 0.6664 mol  $NO$

actual mole ratio = 0.6664 mol  $NO$ /0.5125 mol  $O_2$  = 2.132 mol  $NO$ /1.000 mol  $O_2$

Because the actual mole ratio of  $NO$ / $O_2$  is larger than the balanced equation mole ratio of  $NO$ / $O_2$ , there is an excess of  $NO$ ;  $O_2$  is the limiting reactant.

Mass of  $NO$  used = 0.5125 mol  $O_2$  = 2 mol  $NO$ /1 mol  $O_2$  = 0.4250 mol  $NO$

0.4250 mol  $NO$  = 30.01 g  $NO$ /1 mol  $NO$  = 18.76 g  $NO$

Mass of  $NO_2$  produced = 0.4250 mol  $NO_2$  = 46.01 g  $NO_2$ /1 mol  $NO_2$  = 28.76 g  $NO_2$

Excess  $NO$  = 20.00 g  $NO$  = 18.76 g  $NO$  = 1.24 g  $NO$

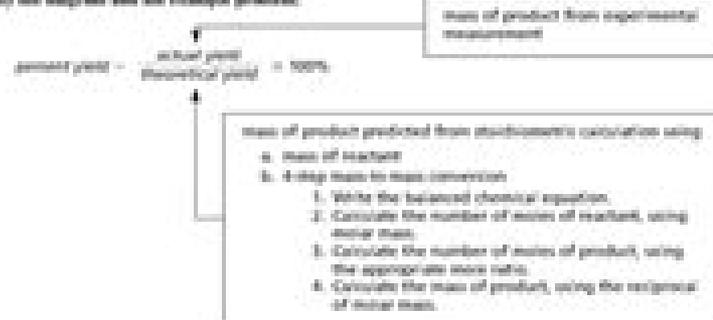
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**CHAPTER 12** **STUDY GUIDE FOR CONTENT MASTERY**

**Section 12.4 Percent Yield**

In your textbook, read about the percent of products.

Study the diagram and the example problem.



**Example Problem:** The following chemical equation represents the production of yellow oxide, a substance used in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the oxide from a 7.00-g sample of yellow oxide. Determine the percent yield of this reaction. The molar masses of  $Ga$  and  $Ga_2O_3$  are 69.72 g/mol and 187.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: unknown 7.42 g  $Ga_2O_2$
- mass of reactant: 7.00 g  $Ga$  correct
- number of moles of reactant: 7.00 g  $Ga$  =  $\frac{69.72 \text{ g } Ga}{1 \text{ mol } Ga}$  7.00 g  $Ga$  = 1 mol  $Ga$ /69.72 g  $Ga$
- number of moles of product: 0.100 mol  $Ga$  =  $\frac{2 \text{ mol } (Ga_2O_2)}{1 \text{ mol } Ga}$  0.100 mol  $Ga$  = 2 mol  $Ga_2O_2$ /4 mol  $Ga$
- theoretical yield: 0.100 mol  $Ga_2O_2$  =  $\frac{187.44 \text{ g } (Ga_2O_2)}{1 \text{ mol } (Ga_2O_2)}$  correct
- percent yield:  $\frac{7.42 \text{ g } (Ga_2O_2)}{7.42 \text{ g } (Ga_2O_2)}$  = 100% 7.42 g  $Ga_2O_2$ /7.42 g  $Ga_2O_2$  = 100

# Standardized Test Practice Chapter 12 Chemistry Answers

**Thomas Griffiths**



**Standardized Test Practice Chapter 12 Chemistry Answers:**

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In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing short of extraordinary. Within the captivating pages of **Standardized Test Practice Chapter 12 Chemistry Answers** a literary masterpiece penned by a renowned author, readers attempt a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting effect on the hearts and minds of those who partake in its reading experience.

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