

## GAS LAWS, GAS STOICHIOMETRY AND KINETIC MOLECULAR THEORY OF GASES

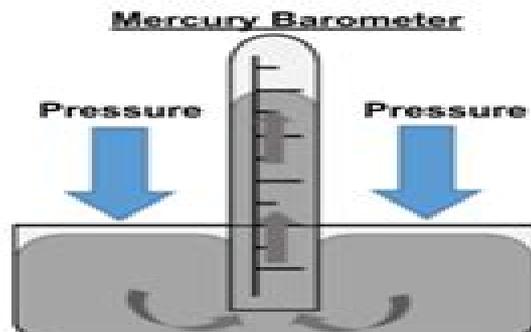
The entire universe is made up of matter including humans, animals, plants, and even the non-living things. There are three phases or states of matter, namely; solid, liquid, and gas. Moreover, gas behaves differently from solids and liquids due to differences in their molecular behavior. The movement of the gas is in random motion due to the vast empty space in a certain system since the distance between particles are far greater than the other two states of matter. Under some conditions of a gas sample, it can be defined in terms of its variables, temperature, volume, number of moles, and pressure.

In the gaseous phase, molecules or the particles collide randomly against other molecules and against their container or their system. This random collision results changes in momentum which give way to the one of the properties of gases called pressure. Pressure, as one of the variables, is defined as the force exerted per unit area. It refers to the force exerted by colliding molecules per unit area of container walls and is mathematically described as  $P = F/A$ .

### Units of Pressure

What makes pressure quantifiable? How is pressure measured and expressed? The following details are some of the units that can be used for pressure:

- Standard Atmosphere (atm)** □ It is commonly and widely used unit for pressure in chemistry. The 1 atm is equal to the pressure that supports a column of mercury which is exactly 760 mmHg.
- Torr (or mmHg)** □ mmHg means millimeter of mercury, where it represents the pressure exerted by a column of a mercury which exactly equals to atmosphere. The unit mmHg is called torr, named after the Italian scientist *Evangelista Torricelli*, who also invented *Barometer*, a device used to measure the atmospheric pressure.
- Pounds per square inch (psi)** □ If we say that gas exerts a pressure of 20 psi, it means the pressure on the wall of the gas container is 20 pounds or the force per square inch of the unit area. The atmospheric pressure at sea level is 14.7 psi.
- kilopascal(kPa)** □ A kilopascal is equal to 1000 pascals (Pa). It is the SI unit for pressure.



### Conversion Factor:

$$1 \text{ atm} = 760 \text{ mmHg} = 760 \text{ torr} = 101.3 \text{ kPa} = 14.7 \text{ psi}$$
$$1 \text{ kPa} = 1000 \text{ Pa}$$

Air is all around us, thus, gases are always involved in our daily activities, from breathing down to automobile or bicycle tires, balloons, and even lifeboats and vests. Life won't be possible without this life-sustaining gas found in the atmosphere.

Scientists have always been curious about how gases behave and how they are different compared to other states of matter. Investigations and experiments on the behavior of gases lead

# Gas Laws And Gas Stiochiometry Study Guide

**Robert K. Wismer**



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## **Table of Contents Gas Laws And Gas Stiochiometry Study Guide**

1. Understanding the eBook Gas Laws And Gas Stiochiometry Study Guide
  - The Rise of Digital Reading Gas Laws And Gas Stiochiometry Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws And Gas Stiochiometry Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Laws And Gas Stiochiometry Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws And Gas Stiochiometry Study Guide
  - Personalized Recommendations
  - Gas Laws And Gas Stiochiometry Study Guide User Reviews and Ratings
  - Gas Laws And Gas Stiochiometry Study Guide and Bestseller Lists
5. Accessing Gas Laws And Gas Stiochiometry Study Guide Free and Paid eBooks
  - Gas Laws And Gas Stiochiometry Study Guide Public Domain eBooks
  - Gas Laws And Gas Stiochiometry Study Guide eBook Subscription Services
  - Gas Laws And Gas Stiochiometry Study Guide Budget-Friendly Options
6. Navigating Gas Laws And Gas Stiochiometry Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Gas Laws And Gas Stiochiometry Study Guide Compatibility with Devices
  - Gas Laws And Gas Stiochiometry Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gas Laws And Gas Stiochiometry Study Guide
  - Highlighting and Note-Taking Gas Laws And Gas Stiochiometry Study Guide
  - Interactive Elements Gas Laws And Gas Stiochiometry Study Guide
8. Staying Engaged with Gas Laws And Gas Stiochiometry Study Guide

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gas Laws And Gas Stiochiometry Study Guide
9. Balancing eBooks and Physical Books Gas Laws And Gas Stiochiometry Study Guide
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gas Laws And Gas Stiochiometry Study Guide
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Gas Laws And Gas Stiochiometry Study Guide
- Setting Reading Goals Gas Laws And Gas Stiochiometry Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Laws And Gas Stiochiometry Study Guide
- Fact-Checking eBook Content of Gas Laws And Gas Stiochiometry Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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