

**Chapter 10 test****Matching**

Match each item with the correct statement below.

- |                            |                                      |
|----------------------------|--------------------------------------|
| a. representative particle | d. percent composition               |
| b. mole                    | e. standard temperature and pressure |
| c. Avogadro's number       | f. empirical formula                 |

- \_\_\_\_\_ 1. the number of representative particles of a substance present in 1 mole of that substance
- \_\_\_\_\_ 2. an atom, an ion, or a molecule, depending upon the way a substance commonly exists
- \_\_\_\_\_ 3. the SI unit used to measure amount of substance
- \_\_\_\_\_ 4. 0°C and 1 atm
- \_\_\_\_\_ 5. the percent by mass of each element in a compound
- \_\_\_\_\_ 6. the smallest whole number ratio of the atoms in a compound

**Multiple Choice**

Identify the letter of the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 7. What SI unit is used to measure the number of representative particles in a substance?
- |             |           |
|-------------|-----------|
| a. kilogram | c. kelvin |
| b. ampere   | d. mole   |
- \_\_\_\_\_ 8. How many hydrogen atoms are in 5 molecules of isopropyl alcohol, C<sub>3</sub>H<sub>7</sub>O?
- |                                     |                                      |
|-------------------------------------|--------------------------------------|
| a. $5 \times (6.02 \times 10^{23})$ | c. 35                                |
| b. 5                                | d. $35 \times (6.02 \times 10^{23})$ |
- \_\_\_\_\_ 9. All of the following are equal to Avogadro's number EXCEPT \_\_\_\_\_
- |  |
|--|
| a. the number of atoms of bromine in 1 mol Br <sub>2</sub>     |
| b. the number of atoms of gold in 1 mol Au                     |
| c. the number of molecules of nitrogen in 1 mol N <sub>2</sub> |
| d. the number of molecules of carbon monoxide in 1 mol CO      |
- \_\_\_\_\_ 10. How many moles of tungsten atoms are in  $4.8 \times 10^{23}$  atoms of tungsten?
- |                               |                               |
|-------------------------------|-------------------------------|
| a. $8.0 \times 10^{-2}$ moles | c. $1.3 \times 10^{-4}$ moles |
| b. $8.0 \times 10^1$ moles    | d. $1.3 \times 10^{-2}$ moles |
- \_\_\_\_\_ 11. How many atoms are in 0.075 mol of titanium?
- |                         |                         |
|-------------------------|-------------------------|
| a. $1.2 \times 10^{23}$ | c. $6.4 \times 10^2$    |
| b. $2.2 \times 10^{24}$ | d. $4.5 \times 10^{22}$ |
- \_\_\_\_\_ 12. How many molecules are in 2.10 mol CO<sub>2</sub>?
- |                                    |                                     |
|------------------------------------|-------------------------------------|
| a. $2.53 \times 10^{24}$ molecules | c. $3.49 \times 10^{-24}$ molecules |
| b. $3.79 \times 10^{24}$ molecules | d. $1.26 \times 10^{24}$ molecules  |

# Chemistry Chapter 10 Test Answers

**Thomas A. Evangelist, Wendy Hanks**



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