

VSA QUESTIONS (1 - MARK QUESTIONS)

1. Define the term 'rate of reaction'.
2. Mention the units of rate of reaction.
3. Express the rate of reaction in terms of $\text{Br}^- (\text{aq})$ as reactant and $\text{Br}_2 (\text{aq})$ as product for the reaction :



4. For a chemical reaction represented by $\text{R} \rightarrow \text{P}$ the rate of reaction is denoted by

$$\frac{-\Delta[\text{R}]}{\Delta t} \text{ or } \frac{+\Delta[\text{P}]}{\Delta t}$$

Why a positive sign (+) is placed before $\frac{\Delta[\text{P}]}{\Delta t}$ and negative sign (-)

before $\frac{\Delta[\text{R}]}{\Delta t}$?

5. Express the rate of reaction in terms of disappearance of hydrogen and appearance of ammonia in the given reaction.



6. Why rate of reaction does not remain constant throughout?
7. Write the unit of first order rate constant of a gaseous reaction if the partial pressure of gaseous reactant is given in bar.
8. For a zero order reaction :

$\text{R} \rightarrow \text{P}$, the change in concentration of reactant w.r.t. time is shown by following graph.

Chemistry Chemical Kinetics Practice Questions Answers

M Walker



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