

Lab 22A

Investigating Chemical Equilibrium

Purpose: To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and relate the the systems, which are exploring the structures through applying Le Chatelier's principle.

Proble:

1. Write a balanced equation for the reaction in part 1.
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
2. Write a balanced equation for the reaction in part 2.
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
3. Write a balanced equation for the reaction in part 3.
 $\text{Co}^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
4. In part 4, what is the conjugate acid of the base CoCl_4^{2-} and what is the conjugate base of the acid CoCl_4^{2-} ?
Conjugate acid: HCoCl_4^{-} (the central group and have the other three not binding). The conjugate base: CoCl_3^{-} (losing three of the groups, the central is the proton and other three must have the other two hydrogen).
5. What is the conjugate acid of the base CoCl_4^{2-} ?
Conjugate acid: HCoCl_4^{-} (the central is the proton and other three not binding).
6. What is the conjugate base of the acid CoCl_4^{2-} ?
Conjugate base: CoCl_3^{-} (losing three of the groups, the central is the proton and other three must have the other two hydrogen).
7. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
 $\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$
8. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
 $\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$
9. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
 $\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$
10. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
 $\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3$

Investigating Chemical Equilibrium Lab 12a Answer

Hussin A.Rothana



Investigating Chemical Equilibrium Lab 12a Answer:

Decoding **Investigating Chemical Equilibrium Lab 12a Answer**: Revealing the Captivating Potential of Verbal Expression

In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Investigating Chemical Equilibrium Lab 12a Answer**," a mesmerizing literary creation penned by a celebrated wordsmith, readers embark on an enlightening odyssey, unraveling the intricate significance of language and its enduring impact on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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