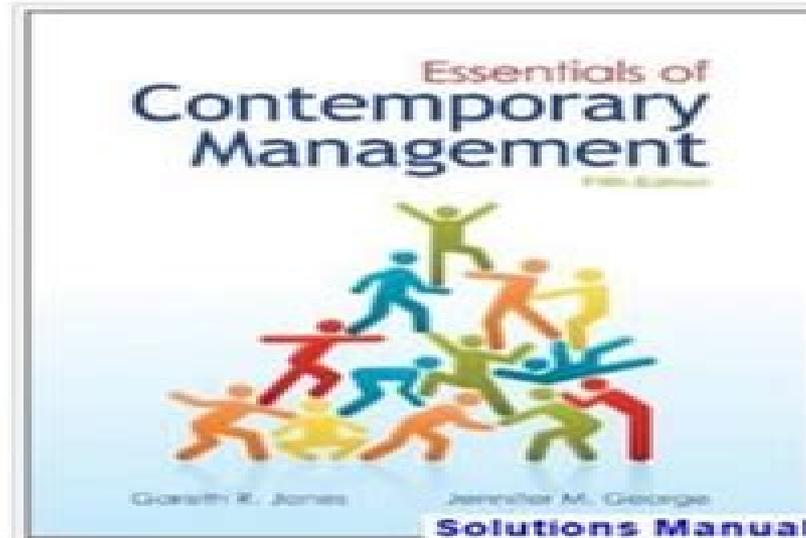


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by its rate law for example the rate of the gas phase decomposition of dinitrogen pentoxide  $2\text{N}_2\text{O}_5 \rightarrow 4\text{NO}_2 + \text{O}_2$  has been found to be directly proportional to the concentration of  $\text{N}_2\text{O}_5$  text rate  $k$  n

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web the initial rate law is then measured for each of the reactions consider the reaction between nitrogen monoxide gas and hydrogen gas to form nitrogen gas and water vapor the following data were collected for this reaction at 1280 text o text c

see table below notice that the starting concentrations of  $\text{ce no}$  and  $\text{ce h}$

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