

# Half Modern Chemistry Review

## CHAPTER 9: STOICHIOMETRY

The following pages contain the best (but not all) of the information for the chapter 9 test.

Focus on this content, but make sure to review class notes, activities, worksheets, questions, etc.

If you study this document and NOTHING else, you should at least be able to PASS the test.

\*\*\*\* Test items will be recall, recognition, and/or application of this content. \*\*\*\*

### OBJECTIVES

- Calculate molar masses for molecular-formula compounds. (C & C)
- Calculate molar masses for ionic compounds. (B & C)
- 9.1. Determine masses of reactants from balanced chemical equations. (T & P)
- 9.2. Perform stoichiometry calculations with: a) mole to mole, mole to gram, gram to mole, and gram to gram. (P & P)
- 9.3. Calculate stoichiometry problems and determine the limiting reagent, excess reactant, and amount of product produced. (P & P)

### 9.1: INTRODUCTION TO STOICHIOMETRY

#### a. Introductory

- composition stoichiometry = calculations involving the mass relationships of elements in compounds.
- reaction stoichiometry = calculations involving the mass relationships between reactants and products in a chemical reaction.
- molar ratio = a conversion factor that relates the amount, in moles, of any two substances involved in a chemical reaction.

#### b. Chapter Highlights

- Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction.
- Stoichiometry is based on the law of conservation of mass.
- Reacting mass is defined as the mass of the substance that reacts.
- A molar ratio is the conversion factor that relates the amount, in moles, of any two substances in a chemical reaction. The molar ratio is derived from the balanced equation.
  - Example:  $2\text{FeCl}_3 + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_4$       (What is the ratio of iron to chlorine moles?) 2:3 or 3:2
- Amount of a substance is measured in moles, and mass of a substance is represented using mass units such as grams, kilograms, or milligrams.
- Mass and amount of substances are quantities, whereas moles and grams are units.
- A balanced chemical equation is necessary to solve any stoichiometric problem.

### 9.2: BASIC STOICHIOMETRIC CALCULATIONS

#### a. Chapter Highlights

- We use molar stoichiometric calculations, the mass, or the amount of any reactant or product over the calculation of the balanced chemical equation, and the mass or amount of any other reactant or product is known.

### 9.3: LIMITED REACTANTS AND PERCENTAGE YIELD

# Modern Chemistry Chapter 9 Section 1 Review Answers

**Siddappa N.Byrareddy**



**Modern Chemistry Chapter 9 Section 1 Review Answers:**

## Unveiling the Power of Verbal Artistry: An Mental Sojourn through **Modern Chemistry Chapter 9 Section 1 Review Answers**

In a global inundated with displays and the cacophony of immediate transmission, the profound power and mental resonance of verbal beauty usually disappear into obscurity, eclipsed by the constant barrage of sound and distractions. Yet, set within the musical pages of **Modern Chemistry Chapter 9 Section 1 Review Answers**, a charming work of fictional elegance that impulses with organic thoughts, lies an unforgettable trip waiting to be embarked upon. Penned by way of a virtuoso wordsmith, that magical opus instructions visitors on an emotional odyssey, softly revealing the latent potential and profound affect embedded within the intricate web of language. Within the heart-wrenching expanse of this evocative evaluation, we shall embark upon an introspective exploration of the book is key subjects, dissect its interesting publishing fashion, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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