

Molality is number of moles of solute present in 1 kg of solvent.

e.g. if 80 g NaOH is present in 1000 g water then the molality will be calculated number of

$$\text{moles of NaOH} = \frac{\text{Amount}}{\text{Molar Mass}}$$

$$= \frac{80}{40} = 2$$

Here 2 moles of NaOH are in 1000 g (or 1 kg)

Solvent so the molality is 2 molal

$$\text{Formulae of molality} = \frac{W_{\text{Solute}}}{M_{\text{Solute}} \times W_{\text{solvent}} (\text{kg})}$$

$$= \frac{W_{\text{Solute}} \times 1000}{M_{\text{Solute}} \times W_{\text{solvent}} (g)}$$

Mol fraction: Fraction of moles of a particular component in mixture.

e.g. In mixture of He(2 mol) and Ne(3 mol) the mol fraction of He = $\frac{2}{2+3} = \frac{2}{5} = 0.4$

the mol fraction of Ne = $\frac{3}{2+3} = \frac{3}{5} = 0.6$

Formula for mol fraction of component A

$$\left[X_A = \frac{n_A}{\sum_{i=1}^n n_i} = \frac{\text{number of moles of A}}{\text{Total number of moles of all species}} \right]$$

Mole Fraction Problems And Solutions

J Ma



Mole Fraction Problems And Solutions:

Unveiling the Power of Verbal Artistry: An Mental Sojourn through **Mole Fraction Problems And Solutions**

In a global inundated with monitors and the cacophony of instantaneous connection, the profound energy and emotional resonance of verbal artistry often diminish in to obscurity, eclipsed by the constant onslaught of sound and distractions. However, located within the lyrical pages of **Mole Fraction Problems And Solutions**, a interesting function of fictional beauty that pulses with fresh feelings, lies an wonderful journey waiting to be embarked upon. Published by way of a virtuoso wordsmith, that enchanting opus instructions viewers on a mental odyssey, lightly exposing the latent potential and profound impact embedded within the complicated web of language. Within the heart-wrenching expanse with this evocative examination, we shall embark upon an introspective exploration of the book is main styles, dissect their interesting publishing model, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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