

Experiment 11A: The Periodic Table

Procedure

Part 1 Electronic Configuration and the Periodic Table

1. Construct a periodic table of your own.

Part 2 Atomic Number, Atomic Radius, Ionization Energy and the Periodic Table

1. Construct a periodic table of your own.
2. Label it as follows.

Part 3 Electronegativity and the Periodic Table

1. Construct a periodic table of your own.

Objectives

1. Explain the placement of elements in groups 1, 2, 10, 11, 12, 13, 14, 15, 16, 17, and 18 on the basis of an element's placement in the periodic table. How does the placement of an element in the periodic table relate to its atomic number?

Explain the placement of elements in groups 1, 2, 10, 11, 12, 13, 14, 15, 16, 17, and 18 on the basis of an element's placement in the periodic table. How does the placement of an element in the periodic table relate to its atomic number?

2. Explain the graph of atomic radius plotted against increasing atomic number (Figure 11A-1). What is particularly noteworthy about it? Is it identical to the periodic table?

Explain the graph of atomic radius plotted against increasing atomic number (Figure 11A-1). What is particularly noteworthy about it? Is it identical to the periodic table? How does the placement of an element in the periodic table relate to its atomic number?

3. Which group of elements has the highest atomic number? Which group has the lowest atomic number? Which group has the highest ionization energy?

Explain the graph of ionization energy plotted against increasing atomic number (Figure 11A-2). What is particularly noteworthy about it? Is it identical to the periodic table? How does the placement of an element in the periodic table relate to its atomic number?

4. Explain the graph of electronegativity plotted against increasing atomic number (Figure 11A-3). What is particularly noteworthy about it? Is it identical to the periodic table? How does the placement of an element in the periodic table relate to its atomic number?

Explain the graph of electronegativity plotted against increasing atomic number (Figure 11A-3). What is particularly noteworthy about it? Is it identical to the periodic table? How does the placement of an element in the periodic table relate to its atomic number?

Laboratory Experiment 11a Answers Periodic Table

Kathleen Armour



Laboratory Experiment 11a Answers Periodic Table:

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