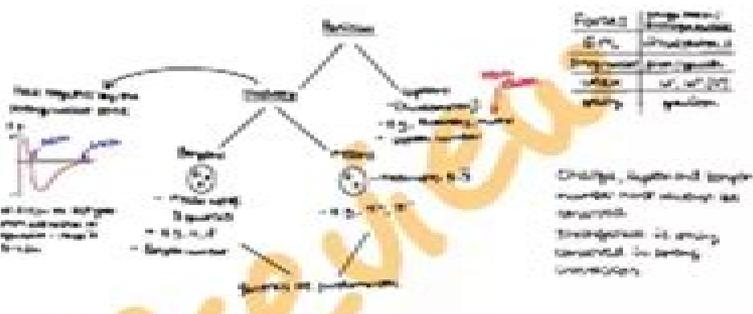
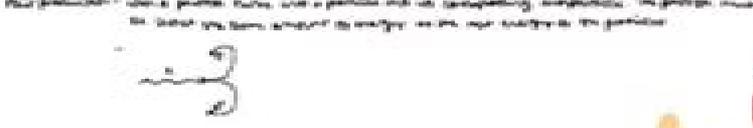


# Particles and Quantum



**Photoelectric effect:** When a photon hits an electron, it can give it energy. If the photon has enough energy, it can knock the electron out of the metal. The energy of the photon is  $E = hf$ .

**Compton effect:** When a photon hits an electron, it can give it energy. The photon must have enough energy to knock the electron out of the metal.



**Excitation:** An atom can be excited by absorbing energy. This energy can come from a photon or from a collision with another atom.

**Alpha particles:** Helium nucleus,  ${}^4_2\text{He}$ . Highly ionizing, so it is dangerous. Stopped by a piece of paper or skin.

**Beta minus:** Fast moving electron,  $e^-$ . Produced from a neutron decaying into a proton and an electron. Stopped by a few cm of aluminium.

**Beta plus:** Fast moving positron,  $e^+$ . Produced from a proton decaying into a neutron and a positron. Stopped by a few cm of aluminium.



**Gamma:** High energy photon. Produced from a high energy nucleus. Highly penetrating, stopped by a few m of concrete.

**Isotopes:** Same element, same number of protons, different number of neutrons.

**Specific charge:**  $\frac{q}{m}$  (C/kg). Charge to mass ratio.

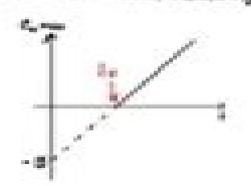
**Electron volt:** Charge of an electron when accelerated through a potential of one volt.  $1 \text{ eV} = 1.6 \times 10^{-19} \text{ J}$ .



## Photoelectric effect

Use photons of sufficient energy on electrons in the surface of a metal, liberating them.

- Work function  $\phi_0$ :** Minimum energy to liberate an electron.
- Stopping potential  $V_s$ :** Minimum potential needed to stop the electrons.
- Maximum kinetic energy  $E_{k,max}$ :**  $E_{k,max} = hf - \phi_0$ .



The photoelectric effect provides the particle nature of light due to the fact that the maximum kinetic energy of electrons is zero for frequencies below a certain value, indicating the energy of light cannot be divided into smaller parts.

- Stopping potential  $V_s$ :** The pot. needed to stop any electrons from reaching the other side.
- Equation:**  $eV_s = hf - \phi_0$ .
- Graphs:**
  - $V_s$  vs  $f$ : Straight line with positive slope.
  - $V_s$  vs  $\lambda$ : Straight line with negative slope.

## De Broglie

The de Broglie wavelength is the wavelength that a particle can have. The wave nature of a particle was proven by diffraction of electrons. This is because the electron is diffracting which is wave-like.



$\lambda = \frac{h}{p}$

$\lambda = \frac{h}{mv}$

$\lambda = \frac{h}{m\gamma v}$

$\lambda = \frac{h}{m\beta c}$

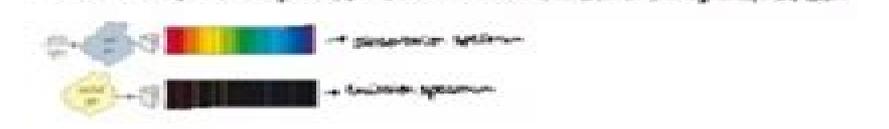
## Fluorescent tubes



These tubes contain mercury vapor. A high voltage is applied. This excites the mercury atoms. When they return to the ground state, they emit light. The phosphor coating is excited by the light and emits visible light.

## Energy levels

Electrons can be excited to higher energy levels. When they return to the ground state, they emit light. The energy of the light is equal to the difference in energy between the two levels.



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**Terry C. Jones**



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