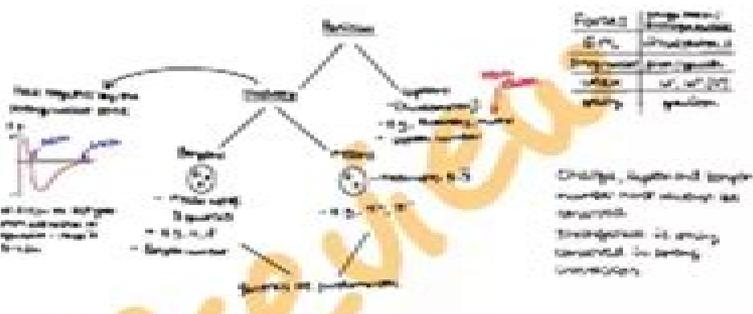


Particles and Quantum



Excitation: when a particle and its corresponding antiparticle collide and their rest energy is converted into photons.



Pair production: when a photon turns into a particle and its corresponding antiparticle. The photon must have at least the sum amount of energy to create the particles.



Excitation

Exciting means they give electrons enough energy to climb their atomic shells.

- Alpha particles - Helium nucleus, ^4_2He
- Highly ionising, so it is dangerous
- Highly penetrating - stopped by a plate of paper or foil in air
- $^{238}\text{U} \rightarrow ^{234}\text{Th} + ^4_2\text{He}$

Beta - fast moving electron $\rightarrow e^-$

- Medium penetrating and ionising ability - stopped by a few cm of aluminium
- Not ionising - $^{14}\text{C} \rightarrow ^{14}\text{N} + e^- + \bar{\nu}_e$ **essential** to avoid superconduct



Gamma - high energy EM ray or photon emitted from a high energy nucleus $\rightarrow \gamma$

- Highly ionising
- Highly penetrating - stopped by a few m of concrete

Isotopes - same element - same number of protons

- different relative atomic mass - different number of neutrons

$$^{12}_6\text{C} + ^{13}_6\text{C}$$

Specific charge - $\frac{Q}{m}$ (C/kg)

Electron rest - charge of an electron when stationary through a bit of air and

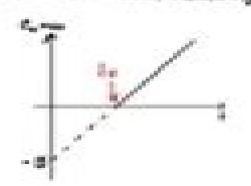
$$-1.6 \times 10^{-19} \text{ C}$$



Photoelectric effect

Use photons of sufficient energy on electrons on the surface of the metal, liberating them.

- $E_{\text{photon}} = hf = \phi + E_{\text{max}}$
- ϕ - photon absorbed - $h \times \text{threshold freq}$ - energy for photon
- f - frequency
- E_{max} - extra photon - maximum energy to liberate electrons
- E_{photon} - energy that will give an e^- its maximum
- ϕ - stopped frequency - minimum frequency needed for electrons to be liberated



The photoelectric effect proved the particle nature of light due to the fact that the maximum kinetic energy of electrons E_{max} is not linearly dependent on the wavelength of light. Instead, it is linearly dependent on the frequency of light. E_{max} is zero at the threshold frequency but is constant at all frequencies above the threshold.

Stopping potential - the pd needed to stop any electrons from reaching the other side.

- $E_{\text{photon}} = hf = \phi + eV_s$
- e - charge of electron
- V_s - stopping potential

De Broglie

The de Broglie wavelength is the wavelength that a particle can have.

The wave nature of a particle will occur by diffraction or a grating from which particles would produce the characteristic pattern. This is because the electron is diffracting which is wave-like.



$$\lambda = \frac{h}{p} = \frac{h}{mv}$$

λ - de Broglie wavelength
 p - momentum / mass x velocity

Fluorescent tubes



This tube contains mercury under a high voltage is applied. This excites the gas resulting in which some of the mercury atoms emit photons when they de-excite.

When the gas de-excites it emits photons in many colors, why the colour is higher energy photons when the electrons de-excite, they emit photons in the UV range. A phosphor coating in the tube absorbs the photons and emit visible light.

Energy levels

Electrons can be excited to higher energy levels. Either by absorbing photons from the light or by the collision. In example below it is a fast electron colliding with a stationary low energy electron.

Ionisation energy - is enough energy to pull an electron out of an atom, so will completely change the atom.



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