

50/50

Experiment 9 Prelaboratory Assignment

A Volumetric Analysis

Date 03/30

Desk No. _____

1. a. What is the titrant for this experiment?

NaOH (base), sodium hydroxide

b. Is the indicator generally added to the titrant or the analyte in a titration?

The indicator, phenolphthalein, is added to the analyte, the acid.
It is colorless in an acid but pink in basic solutions.

2. a. What is the primary standard used in this experiment (name and formula)? Define a primary standard.

A primary standard is an accurate mass of a solid substance which is measured, dissolved in water, and then reacted.

↳ $\text{KHC}_8\text{H}_4\text{O}_4$, potassium hydrogen phthalate

b. What is the secondary standard used in this experiment (name and formula)? Define a secondary standard.

A secondary standard is the standard remaining solution from after the molar concentration of the titrant is measured.

↳ NaOH, sodium hydroxide

3. Distinguish between a stoichiometric point and an endpoint in an acid-base titration.

reaction is complete
because amount
reactants that have
combined

when the indicator changes
color due to the change
in pH

4. a. When rinsing a buret after cleaning it with soap and water, should the rinse be dispensed through the buret tip or the top opening of the buret? Explain.

The water and rinse should go through the tip. Solution will dispense through and help keep the concentration of the titrant the same.



b. Experimental Procedure, Part A.4. In preparing the buret for titration, the final rinse is with the NaOH titrant rather than with deionized water. Explain.

The water can alter the values and the NaOH is hygroscopic.

5. Part A.3. Calculate the mass of $\text{KHC}_8\text{H}_4\text{O}_4$ (molar mass = 204.23 g/mol) that reacts with 15 mL of the 0.15 M NaOH solution prepared in Part A.2. Express this mass $\text{KHC}_8\text{H}_4\text{O}_4$ to the correct number of significant figures and record the calculation on the Report Sheet.

$$15 \text{ mL} = 0.015 \text{ L}$$

$$0.015 \text{ L (0.15 M)} = 0.00225 \text{ mol NaOH}$$

$$0.00225 \text{ mol NaOH} \cdot \frac{1 \text{ mol KHC}_8\text{H}_4\text{O}_4}{1 \text{ mol NaOH}} = 0.00225 \text{ mol KHC}_8\text{H}_4\text{O}_4$$

$$0.00225 \text{ mol KHC}_8\text{H}_4\text{O}_4 \cdot \left(\frac{204.23 \text{ g}}{1 \text{ mol}} \right) = \boxed{0.4595 \text{ g}}$$

Chem File Experiment A5 Answers

Brendan G. Carr



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1. Understanding the eBook Chem File Experiment A5 Answers
 - The Rise of Digital Reading Chem File Experiment A5 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem File Experiment A5 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chem File Experiment A5 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem File Experiment A5 Answers
 - Personalized Recommendations
 - Chem File Experiment A5 Answers User Reviews and Ratings

- Chem File Experiment A5 Answers and Bestseller Lists
- 5. Accessing Chem File Experiment A5 Answers Free and Paid eBooks
 - Chem File Experiment A5 Answers Public Domain eBooks
 - Chem File Experiment A5 Answers eBook Subscription Services
 - Chem File Experiment A5 Answers Budget-Friendly Options
- 6. Navigating Chem File Experiment A5 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chem File Experiment A5 Answers Compatibility with Devices
 - Chem File Experiment A5 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chem File Experiment A5 Answers
 - Highlighting and Note-Taking Chem File Experiment A5 Answers
 - Interactive Elements Chem File Experiment A5 Answers
- 8. Staying Engaged with Chem File Experiment A5 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chem File Experiment A5 Answers
- 9. Balancing eBooks and Physical Books Chem File Experiment A5 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem File Experiment A5 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chem File Experiment A5 Answers
 - Setting Reading Goals Chem File Experiment A5 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chem File Experiment A5 Answers
 - Fact-Checking eBook Content of Chem File Experiment A5 Answers
 - Distinguishing Credible Sources

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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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