

Chemistry Semester 2 Final Exam Study Guide

Name _____

- ✓ The final exam will be all multiple-choice. Come prepared with a pencil.
- ✓ Do not bring electronic devices to the exam (other than a calculator). This includes cell phones and mp3 players. If I see them I will take them.
- ✓ There will be no rest room passes until you are done with the test.
- ✓ I will provide a formula sheet (identical to one attached to this study guide) and a laminated periodic table.
- ✓ This study guide will be extra credit toward your final exam grade.
- ✓ **BRING A PENCIL AND A CALCULATOR!!!** Don't come to my final without these.
- ✓ The Chemistry Final Exam will count for 15% of your final semester grade. The exam will cover material from second semester. Some questions will require you to know material from first semester. Complete this review sheet on a separate sheet of paper. This review sheet covers all of the questions you will answer on the final exam, but be prepared to apply the knowledge on this study guide to other questions.

"Success is a state of mind. If you want success, start thinking of yourself as a success."

- ✓ Keys to success:
 - Complete this study guide.
 - Don't wait until Monday or Tuesday before finals to get help.
 - Start studying early. Remember, you have all summer to relax.
 - Be prepared.

Final exam topics:

1. Chemical Reactions + Balancing
2. The Mole + Stoichiometry
3. Aqueous solutions and Solubility
4. The Behavior of Gases
5. Acids and Bases

Final Exam Schedule

Find your class period and time. **DO NOT** be late. You will not be compensated for time you are late and your name will be turned into the office. There is **NO** excuse to miss the final exam.

Thursday, May 27

8:10-9:40	Period 3
9:50-11:20	Period 4
11:20-12:15	Lunch
12:20-1:50	Period 5
2:00-3:30	Period 7

Friday, May 28

8:10-9:40	Period 1
9:50-11:20	Period 2
11:20-12:15	Lunch
12:20-1:50	Period 6
2:00-3:30	Period 8 Make ups

Chemistry Semester 2 Final Study Guide Answers

Christian G. Meyer



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web tersinmez için $w = q - 12.4 \text{ kJ}$ $o = 4.50 \text{ atm}$ basınç altında ve 180 c sıcaklıkta ideal bir gaz $c = 1.5 \text{ r}$ adyabatik şartlarda v a tersinir olarak b sabit bir dış basınca karşı 1 atm e kadar genişliyor