

CHEMISTRY 9TH - Solved Numerical

→ FUNDAMENTALS OF CHEMISTRY

CHAPTER# 01

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1. Calculate the number of moles in 30 gm of H_3PO_4 .

Solution:

Data: Mass of H_3PO_4 = 30 gm

$$\begin{aligned} \text{Molecular mass of } H_3PO_4 &= 1(3) + 30.97 + 16(4) \\ &= 3 + 30.97 + 64 \\ &= 97.97 \text{ or } 98 \end{aligned}$$

Calculation:

$$\text{Number of Moles} = \frac{30}{98} = \boxed{0.3061 \text{ moles}} \text{ Ans}$$

2. Calculate the number of moles, in 50 g of each:

(i) Na (ii) H_2O

$$\text{Number of Moles} = \frac{\text{Given mass of substance}}{\text{Atomic mass or Formula mass}}$$

(i) Given,

(i) Number of moles = ?

(ii) Given mass of Na = 50 g

(iii) Atomic mass of Na = 23 a.m.u.

$$\Rightarrow \text{No of moles of Na} = \frac{50}{23} = 2.173 \text{ moles (of Na)}$$

(ii) Given,

(i) Number of moles of H_2O = ?

(ii) Formula mass of H_2O = 18 a.m.u.

(iii) Given mass of H_2O = 50 g

$$= \text{No of moles of } H_2O = \frac{50}{18} = 2.777 \text{ moles (of } H_2O)$$

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Arshad Iqbal



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