

Key

Predicting Products of Chemical Reactions

This worksheet is designed to help you predict products of simple reactions of the four basic reaction types (synthesis, decomposition, single replacement, and double replacement) and combustion reactions.

For the first few reactions, the type of reaction is listed, you should predict the products, then balance. Further questions just have the reactants listed and you should decide on the type of reaction, as well as the correct products. Many of these reactions fall into the category of redox reactions, though do not let that confuse you...each can be described in terms of the four basic reaction types (except the combustion reactions).

Although states (s, l, g, aq) of the reactants and products are very important in a chemical reaction, don't worry about determining those for these problems. Rather, focus on what products might result from the reactants given. Pay particular attention to the ionic charge of species that you know forms ions with only one possible charge (e.g., alkali metals, alkaline earth metals, halogens, etc.)

1. Combustion: $C_2H_6 + 9 O_2 \rightarrow 6 CO_2 + 6 H_2O$
2. Combustion: $2 C_2H_6 + 11 O_2 \rightarrow 8 CO_2 + 6 H_2O$
3. Combustion: $C_2H_6O_2 + 7 O_2 \rightarrow 6 CO_2 + 5 H_2O$
4. Synthesis: $Mg + I_2 \rightarrow MgI_2$
5. Double displacement: $CuCl_2 + H_2S \rightarrow CuS + 2 HCl$
6. Double displacement: $NaOH + HClO_4 \rightarrow NaClO_4 + H_2O$
7. Decomposition: $ZnCO_3 + \text{heat} \rightarrow ZnO + CO_2$
8. Single replacement: $2 HCl + Zn \rightarrow ZnCl_2 + H_2$
9. single $2 Na + MgCl_2 \rightarrow 2 NaCl + Mg$
10. double $CaCl_2 + K_2CO_3 \rightarrow CaCO_3 + 2 KCl$
11. synthesis $2 K + Cl_2 \rightarrow 2 KCl$
12. double $3 BaCl_2 + 2 K_3PO_4 \rightarrow Ba_3(PO_4)_2 + 6 KCl$
13. double $H_2SO_4 + 2 KOH \rightarrow K_2SO_4 + 2 H_2O$
14. Synthesis $4 Al + 3 O_2 \rightarrow 2 Al_2O_3$
15. double $Pb(NO_3)_2 + 2 KOH \rightarrow Pb(OH)_2 + 2 KNO_3$
16. double $H_2SO_4 + BaCl_2 \rightarrow 2 HCl + BaSO_4$
17. single $Ca + 2 AgCl \rightarrow CaCl_2 + 2 Ag$
18. double $H_3PO_4 + FeBr_3 \rightarrow 3 HBr + FePO_4$
19. Synthesis $6 Li + N_2 \rightarrow 2 Li_3N$
20. double $2 HCl + Mg(OH)_2 \rightarrow MgCl_2 + 2 H_2O$

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Rachel S Tattersall



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Table of Contents Chapter 10 Supplemental Problems Chemical Reactions Answer Key

1. Understanding the eBook Chapter 10 Supplemental Problems Chemical Reactions Answer Key
 - The Rise of Digital Reading Chapter 10 Supplemental Problems Chemical Reactions Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 10 Supplemental Problems Chemical Reactions Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 10 Supplemental Problems Chemical Reactions Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 10 Supplemental Problems Chemical Reactions Answer Key

Chapter 10 Supplemental Problems Chemical Reactions Answer Key

- Personalized Recommendations
 - Chapter 10 Supplemental Problems Chemical Reactions Answer Key User Reviews and Ratings
 - Chapter 10 Supplemental Problems Chemical Reactions Answer Key and Bestseller Lists
5. Accessing Chapter 10 Supplemental Problems Chemical Reactions Answer Key Free and Paid eBooks
- Chapter 10 Supplemental Problems Chemical Reactions Answer Key Public Domain eBooks
 - Chapter 10 Supplemental Problems Chemical Reactions Answer Key eBook Subscription Services
 - Chapter 10 Supplemental Problems Chemical Reactions Answer Key Budget-Friendly Options
6. Navigating Chapter 10 Supplemental Problems Chemical Reactions Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - Chapter 10 Supplemental Problems Chemical Reactions Answer Key Compatibility with Devices
 - Chapter 10 Supplemental Problems Chemical Reactions Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chapter 10 Supplemental Problems Chemical Reactions Answer Key
 - Highlighting and Note-Taking Chapter 10 Supplemental Problems Chemical Reactions Answer Key
 - Interactive Elements Chapter 10 Supplemental Problems Chemical Reactions Answer Key
8. Staying Engaged with Chapter 10 Supplemental Problems Chemical Reactions Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 10 Supplemental Problems Chemical Reactions Answer Key
9. Balancing eBooks and Physical Books Chapter 10 Supplemental Problems Chemical Reactions Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 10 Supplemental Problems Chemical Reactions Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 10 Supplemental Problems Chemical Reactions Answer Key
- Setting Reading Goals Chapter 10 Supplemental Problems Chemical Reactions Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 10 Supplemental Problems Chemical Reactions Answer Key

- Fact-Checking eBook Content of Chapter 10 Supplemental Problems Chemical Reactions Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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