

**CHAPTER 14 REVIEW***Acids and Bases***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. Name the following compounds as acids:

_____ sulfuric acid	a. $\text{H}_2\text{SO}_4$
_____ sulfurous acid	b. $\text{H}_2\text{SO}_3$
_____ hydrosulfuric acid	c. $\text{H}_2\text{S}$
_____ perchloric acid	d. $\text{HClO}_4$
_____ hydrocyanic acid	e. hydrogen cyanide

2. \_\_\_\_\_
- $\text{H}_2\text{S}$
- Which (if any) of the acids mentioned in item 1 are binary acids?

3. Write formulas for the following acids:

_____ $\text{HNO}_2$	a. nitrous acid
_____ $\text{HBr}$	b. hydrobromic acid
_____ $\text{H}_3\text{PO}_4$	c. phosphoric acid
_____ $\text{CH}_3\text{COOH}$	d. acetic acid
_____ $\text{HClO}$	e. hypochlorous acid

4. Calcium selenate has the formula
- $\text{CaSeO}_4$
- .

_____ $\text{H}_2\text{SeO}_4$	a. What is the formula for selenic acid?
_____ $\text{H}_2\text{SeO}_3$	b. What is the formula for selenous acid?

5. Use an activity series to identify two metals that will not generate hydrogen gas when treated with an acid.

Choose from Cu, Ag, Au, Pt, Pd, or Hg.  
\_\_\_\_\_

6. Write balanced chemical equations for the following reactions of acids and bases:

- a. aluminum metal with dilute nitric acid



- b. calcium hydroxide solution with acetic acid



# Chapter 14 Review Acids Bases Mixed Answers

**L Towne**



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