

17

WATER AND AQUEOUS SYSTEMS

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SECTION 17.1 LIQUID WATER AND ITS PROPERTIES (pages 475–478) ◆◆◆

This section explains the hydrogen bonding that occurs in water molecules and how high surface tension and low vapor pressure can be explained in terms of hydrogen bonding.

► The Water Molecule (pages 475–476)

- What unique molecular compound is the foundation of all life on Earth?
water
- Approximately what fraction of Earth's surface is covered in water? 3/4
- Circle the letter next to each sentence that is true concerning water molecules.
 - Each O — H covalent bond in a water molecule is nonpolar.
 - In a water molecule, the less electronegative hydrogen atoms acquire a slight positive charge and the oxygen atom acquires a slight negative charge.
 - Because the water molecule has an H — O — H bond angle of 105°, the molecule as a whole is polar.
- The diagram below depicts a water molecule. Complete the labels showing the locations of the hydrogen atoms, the oxygen atom, and the regions of positive and negative charge.



- List five important properties of water that result from hydrogen bonding.
 - high surface tension
 - low vapor pressure
 - high specific heat capacity
 - high heat of vaporization
 - high boiling point

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