

14 • The Chemistry of Acids & Bases**STUDY QUESTIONS & PROBLEMS - KEY**

1. For the following aqueous equilibria, designate the Bronsted-Lowry conjugate acid-base pairs and establish the weaker side:



2. Write the name and formula for the conjugate bases of the following:

- a. HNO_2 NO_2^- nitrite
 b. H_2SO_4 HSO_4^- bisulfate
 c. H_2PO_4^- HPO_4^{2-} dihydrogen phosphate
 d. HF F^- fluoride
 e. $\text{CH}_3\text{CO}_2\text{H}$ CH_3COO^- acetate

3. Complete the Bronsted-Lowry equilibria, label the components acid or base, and pair up the conjugate acid-base pairs:



4. Is the monohydrogen phosphate ion HPO_4^{2-} amphiprotic? — YES
 If so, write the formulas of its conjugate acid and its conjugate base. CA = H_2PO_4^- and CB = PO_4^{3-}

Chapter 17 Study Acids Bases

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