

Chapter 10 Guided Notes

Name \_\_\_\_\_

1. \_\_\_\_\_ attractions are bonds that hold compounds together
  - a. What are 2 examples of these attractions?
2. \_\_\_\_\_ attractions are between molecules.
3. Which attraction is stronger?
4. Which state of matter will have the **strongest** intermolecular forces? Which state has the **weakest**?



5. Order the intermolecular forces (IMF) in order of weakest to strongest: hydrogen bonding, dipole-dipole, ion-dipole, London dispersion
6. List all the possible IMF for the following compounds:
  - a. HCl
  - b. NH<sub>3</sub>
  - c. H<sub>2</sub>O
  - d. CH<sub>4</sub>
  - e. NaCl dissolved in water

7. For each type of compound determine what forces are possible:

| Type of IMF       | Atoms | Nongpolar molecules | Polar molecules without OH, NH, or HF groups | Polar molecules containing OH, NH, or HF groups | Ionic solids dissolved in polar liquids |
|-------------------|-------|---------------------|--|---|---|
| Dispersion forces |       |                     |  |   |   |
| Dipole-dipole     |       |                     |  |   |   |
| Hydrogen bonding  |       |                     |  |   |   |
| Ion-dipole        |       |                     |  |   |   |

Explore the following procedure and complete the subsequent practice exercises in preparation for the synchronous class discussion. Step 1 or Step 2 are review from general chemistry I. Review Lewis structures

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**Brendan G. Carr**



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