

## Chapter 16 Solubility and Complex Ion Equilibria

### Section 16.1: Solubility Equilibria and the Solubility Product

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#### ◆ 1. What is Solubility Equilibrium?

- When an **ionic solid** (like AgCl) is put in water, some dissolves into ions.
- At the same time, ions in the water can recombine and form the solid again.
- Eventually, it reaches **equilibrium**: the rate of dissolving = rate of crystallizing.



#### ◆ 2. Solubility Product (K<sub>sp</sub>)

- **K<sub>sp</sub>** is the equilibrium constant for this dissolving.
- For AgCl:  $K_{sp} = [Ag^+][Cl^-]$

👉 Notice: The solid (AgCl) is left out because its concentration is constant.

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#### ◆ 3. Solubility vs. K<sub>sp</sub>

- **Solubility** = how many moles per liter dissolve.
- **K<sub>sp</sub>** = a number from the ion concentrations at equilibrium.

⚠ Important: A larger K<sub>sp</sub> doesn't always mean higher solubility because the ion ratios matter.

Example:

- AgCl:  $K_{sp} = 1.6 \times 10^{-10}$ , solubility  $\approx 1.3 \times 10^{-5}$ .
  - Ag<sub>3</sub>PO<sub>4</sub>:  $K_{sp} = 1.8 \times 10^{-18}$ , solubility  $\approx 1.6 \times 10^{-5}$ .  
Even though Ag<sub>3</sub>PO<sub>4</sub> has a smaller K<sub>sp</sub>, its solubility is close because of the 3:1 ratio of ions.
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