

Section 12.1

Section 12.1

Answer complete answer multiple-choice questions that are provided.

1. _____
 - a. _____
 - (a) correct, in terms of all four parts and products.
 - (b) correct, in terms of number of reactants and products.
 - (c) correct, in terms of work done on work done by system.
 - (d) correct, in terms of enthalpy change in the reaction.
 - b. _____
 - (a) the mass number of all the reactants combined is equal to the mass number of all the products combined.
 - (b) the mass of carbon produced is equal to the mass of carbon consumed.
 - (c) the number of moles of hydrogen that will react with a certain quantity of oxygen.
 - (d) the amount of energy required to heat the water from 25°C.
 - c. _____
 - (a) balanced chemical equation always use to describe the reaction.
 - (b) number of moles of reactants in the reaction.
 - (c) energy released in the reaction.
 - (d) relative enthalpies of all elements in the reaction.
 - (e) number of molecules involved in the reaction.
 - d. _____
 - (a) relative number of moles of hydrogen in water in oxygen molecule in water.
 - (b) number of moles of water.
 - (c) number of moles of oxygen.
 - (d) number of moles of hydrogen.
 - (e) number of moles of oxygen.

2. Given the reaction represented by the following balanced equation:



a. Balance the equation:



b. What is the coefficient of H_2O in H_2O ?

c. How many moles of H_2O_2 are formed from 1.0 mole of O_2 in the reaction?

d. How many moles of H_2O are needed to form 1.0 mole of H_2O_2 in the reaction?

e. How many grams of H_2O_2 are formed from 1.0 mole of O_2 in the reaction?

Chapter 9 Stoichiometry Section 1 Answers

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Chapter 9 Stoichiometry Section 1 Answers:

Decoding **Chapter 9 Stoichiometry Section 1 Answers**: Revealing the Captivating Potential of Verbal Expression

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