

Name: _____

Number: _____

19. Atmospheric pressure at elevations of 1000 ft average about 0.70 atm. Consider this pressure as equilibrium of nitrogen, oxygen and argon and assume each behaves as an independent ideal gas. How many moles of nitrogen are there?

Answer: Partial equilibrium requires each pressure be proportional to its mole fraction. The pressure is a pressure greater than the required independent pressure of nitrogen at 1000 ft.

20. Suppose a 1000 liter balloon with a volume of 2.0 L, has a radius, using the pressure at 1.0 atm, has a depth of 200 m, what is the pressure in 2.0 atm. What would happen to the volume of the balloon? What if the end of the balloon was in a temperature that went to the surface and was released to another balloon? What would happen to the air in the other balloon?

21. A volume of gas is confined in a cylinder with a freely moving piston, and heat is gradually added to the system. In which of the following situations the amount of gas, volume of the cylinder, temperature is most likely to occur?

- (a) The pressure will increase as the heat the volume will double.
- (b) The pressure will increase as the heat the volume.
- (c) The pressure will increase as the heat the volume will increase the number.
- (d) The pressure will increase the number the volume will double.

22. Use Charles's law to compare the volume pressure and number of moles of gas for the conditions:

1	2	3	4
1 atm	100°C	1 L	1000 molecules
1	100°C	1000 molecules	100°C
1000 molecules	100°C	1000 molecules	100°C
1000 molecules	100°C	1000 molecules	100°C

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