

Name _____ Date _____ Class _____

CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

States of Matter

Section 13.1 Gases

In your textbook, read about the kinetic molecular theory.

Complete each statement.

- The kinetic molecular theory describes the behavior of gases in terms of particles in motion.
- The kinetic molecular theory makes the following assumptions:
 - In a sample of a gas, the volume of the gas particles themselves is very small compared to the volume of the sample.
 - Between gas particles and the walls, there are no significant attractive or repulsive forces between gas particles.
 - Gas particles are in constant and random motion.
 - The collisions between gas particles are elastic, that is, no kinetic energy is lost.
- The kinetic energy of a particle is represented by the equation $KE = \frac{1}{2}mv^2$.
- Temperature is a measure of the average kinetic energy of the particles in a sample of matter.

In your textbook, read about explaining the behavior of gases.

For each statement below, write true or false.

- | | |
|--------------|----------------------------------------------------------------------------------------------------------------------------------|
| <u>True</u> | 1. Gases are less dense than solids because there is a lot of space between the particles of a gas. |
| <u>True</u> | 2. The random motion of gas particles causes a gas to expand until it fills its container. |
| <u>False</u> | 3. The density of a gas decreases as it is compressed. |
| <u>True</u> | 4. A gas can flow into a space occupied by another gas. |
| <u>True</u> | 5. The diffusion of a gas is caused by the random motion of the particles of the gas. |
| <u>False</u> | 6. Lighter gas particles diffuse less rapidly than do heavier gas particles. |
| <u>True</u> | 7. During effusion, a gas escapes through a tiny opening into a vacuum. |
| <u>False</u> | 8. Graham's law of effusion states that the rate of effusion for a gas is directly related to the square root of its molar mass. |

Study Guide for Content Mastery

Chemistry: Matter and Change • Chapter 13 79

Name _____ Date _____ Class _____

CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

Section 13.1 continued

In your textbook, read about gas pressure.

Circle the letter of the choice that best completes the statement or answers the question.

- Pressure is defined as force per unit
 - area
 - time
 - mass
 - volume
- What is an instrument designed to measure atmospheric pressure?
 - barometer
 - manometer
 - physoconimeter
 - densimeter
- The height of the liquid in a barometer is affected by all of the following EXCEPT the
 - altitude
 - density of the liquid in the column
 - atmospheric pressure
 - diameter of the column tube.
- The pressure of the gas in a manometer is directly related to which of the following
 - height of the mercury column in the closed-end arm
 - height of the mercury column in the open-end arm
 - $h + 76$
 - $h - 76$
- One atmosphere is equal to a pressure of
 - 76 mm Hg
 - 101.3 kPa
 - 147 psi
 - 760 mm
- The partial pressure of a gas depends on all of the following EXCEPT the
 - concentration of the gas
 - size of the container
 - density of the gas
 - temperature of the gas
- The pressure of a sample of air in a manometer is 101.3 kPa. What is the partial pressure of nitrogen (N_2) in the sample if the combined partial pressures of the other gases is 22.4 kPa?
 - 42.4 kPa
 - 78.9 kPa
 - 101.3 kPa
 - 123.7 kPa

Use the figure to answer the following questions.

- What instrument is illustrated in the figure? barometer
- Who invented this instrument? Evangelista Torricelli
- What are the two opposing forces that control the height of the mercury in the column? gravity and atmospheric pressure
- What does it mean when the level of mercury drops in the column? The atmospheric pressure has increased.



80 Chemistry: Matter and Change • Chapter 13

Study Guide for Content Mastery

Chemistry Matter And Change Chapter 7 Study Guide Answers

Puja Mehta



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Table of Contents Chemistry Matter And Change Chapter 7 Study Guide Answers

1. Understanding the eBook Chemistry Matter And Change Chapter 7 Study Guide Answers
 - The Rise of Digital Reading Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Matter And Change Chapter 7 Study Guide Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Personalized Recommendations
 - Chemistry Matter And Change Chapter 7 Study Guide Answers User Reviews and Ratings
 - Chemistry Matter And Change Chapter 7 Study Guide Answers and Bestseller Lists
5. Accessing Chemistry Matter And Change Chapter 7 Study Guide Answers Free and Paid eBooks
 - Chemistry Matter And Change Chapter 7 Study Guide Answers Public Domain eBooks
 - Chemistry Matter And Change Chapter 7 Study Guide Answers eBook Subscription Services
 - Chemistry Matter And Change Chapter 7 Study Guide Answers Budget-Friendly Options
6. Navigating Chemistry Matter And Change Chapter 7 Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Matter And Change Chapter 7 Study Guide Answers Compatibility with Devices
 - Chemistry Matter And Change Chapter 7 Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Highlighting and Note-Taking Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Interactive Elements Chemistry Matter And Change Chapter 7 Study Guide Answers
8. Staying Engaged with Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Matter And Change Chapter 7 Study Guide Answers
9. Balancing eBooks and Physical Books Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Matter And Change Chapter 7 Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Setting Reading Goals Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Matter And Change Chapter 7 Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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